

Unit 7: Reduction, Oxidation and Electrochemistry

Chapter 17: Electrochemistry

4.9: Oxidation-Reduction Reactions

Reduction-Oxidation Reactions (Redox Rxn): - chemical reactions where there is a transfer of electron(s).

Oxidation States (Oxidation Number): - a number that is arbitrary assigned to an atom in an element, molecule, and polyatomic ions to account for the number of electrons in *redox* reaction.

Rules for Assigning Oxidation Numbers:

1. **Natural Elements (include diatomic and polyatomics) have an Oxidation Number of 0.**

Examples: Na_(s), O_{2(g)}, O_{3(g)}, H_{2(g)}, F_{2(g)}, P_{4(s)}, and Hg_(l) all have an oxidation number of 0.

2. **Single Atomic Ions have an Oxidation Number Equals to its Charge.**

Example: K⁺ has an oxidation number of +1.

3. **Oxygen in Binary Compound and Polyatomic Ions has an Oxidation Number of -2.**

Examples: CO₂, CO, SO₃, SO₂, CO₃²⁻, SO₄²⁻ and H₂O all have an oxidation number of -2 for oxygen.

An exception occurs in peroxides (compound containing O₂²⁻) where the oxygen has an oxidation number of -1.

4. **Hydrogen in Binary Compounds and Polyatomic Ions has an Oxidation Number of +1.**

Examples: H₂O, HCl, HBr, CH₄, NH₄⁺ and H₂S all has an oxidation number of +1 for hydrogen.

5. **In Binary Compounds, the Atom that has the most attraction to electrons (HIGHER Electronegativity) is assigned the NEGATIVE Oxidation Number. This Negative Oxidation Number is the SAME as its Ionic Charge.**

Examples: HF: F is more electronegative. Thus, F has an oxidation number of -1.

NH₃: N is more electronegative. Thus, N has an oxidation number of -3.

H₂S: S is more electronegative. Thus, S has an oxidation number of -2.

6. **The Sum of all Oxidation Numbers is 0 for all Electrically Neutral Compounds.**

Examples: CO: Oxidation numbers: O = -2; C = +2 [(+2) + (-2) = 0]

SO₂: Oxidation numbers: O = -2; S = +4 [(+4) + 2(-2) = 0]

7. **The Sum of all Oxidation Numbers is Equal to the Overall Charge for all Polyatomic Ions.**

Examples: CO₃²⁻: Oxidation numbers: O = -2; C = +4 [(+4) + 3(-2) = -2]

SO₄²⁻: Oxidation numbers: O = -2; S = +6 [(+6) + 4(-2) = -2]

Note: Same Atom from different chemical species might NOT have the Same Oxidation Number.

We write the charges of ionic species as *n*⁺ or *n*⁻. Oxidation numbers are written as +*n* or -*n*.

Oxidation Reduction Study Guide Answers

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