

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 2-? _____ (anion)
which cation of groups 1 or 2 forms a _____ (cation).
2. An element in the halogen group has _____ valence electrons, and needs to _____ 1 more to have 8. This is because it has _____ valence.
3. An _____ bond is the attraction between oppositely charged ions. HCl is a chemical bond is _____.
4. An ionic bond is formed between _____, which is attracted by an ionic bond, whereas _____, which is not.
5. Elements in the _____ group always form ions with a charge of 2- as an anion. Bonded to _____, and always forms an ionic bond. This is because it has _____ valence electrons. It is _____, and it needs to _____ 2 more to have 8 valence electrons. It is _____.
6. An ionic bond is formed between _____, which is attracted by an ionic bond, whereas _____, which is not.
7. Elements in the _____ group always form ions with a charge of 2- as an anion. Bonded to _____, and always forms an ionic bond.

B. Formulas for Ionic Compounds - Fill in the blank below

Write the correct chemical formula for the compound.	Aluminum	Barium	Potassium	Lithium
Chlorine				
Oxygen				

C. Write the chemical formula for each of the following ions:

Write your answer below the Lewis structure, and use arrows to show the transfer of electrons. Use the periodic table to find the oxidation number. They will combine so that the oxidation numbers cancel out to zero. Use the criss cross method if the oxidation numbers do not cancel.

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S Marginson



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