

SOLUTION

$$w = - \int_{V_1}^{V_2} P dV \quad P = \frac{RT}{V-b} - \frac{a}{V^2}$$

$$w = - \int_{V_1}^{V_2} \frac{RT}{v-b} dV \quad \text{?} \quad \int_{V_1}^{V_2} \frac{a}{V^2} dV$$

$$= - RT \ln \frac{V_2-b}{V_1-b} + a \left( \frac{1}{V_1} - \frac{1}{V_2} \right)$$

✓1.9 Are the following expressions exact differentials?

(a)  $xy^2 dx - x^2 y dy$

(b)  $\frac{dx}{y} - \frac{x}{y^2} dy$

SOLUTION

(a) Taking the cross derivatives

$$\frac{\partial xy^2}{\partial y} = 2xy$$

$$\frac{\partial (-x^2 y)}{\partial x} = -2xy$$

Since the cross derivatives are unequal, the expression is not an exact differential.

(b)  $\frac{\partial (1/y)}{\partial y} = - \frac{1}{y^2}$

$$\frac{\partial (-x/y^2)}{\partial x} = - \frac{1}{y^2}$$

Since the cross derivatives are equal, the expression is an exact differential.

1.10 A mole of liquid water is vaporized at 100 °C and 1.013 bar. The heat of vaporization is 40.69 kJ mol<sup>-1</sup>. What are the values of (a)  $w_{rev}$ , (b)  $q$ , (c)  $\Delta U$ , and (d)  $\Delta H$ ?

SOLUTION

(a) Assuming that water vapor is a perfect gas and that the volume of liquid water

# Physical Chemistry Robert Alberty Solution Manual

**Jessica J Manson**



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