

Chemical Nomenclature Review

Introduction

This is a worksheet reviewing some of the important concepts you will need to use from previous chapters to be successful with stoichiometry. Each section has a sample problem followed by a series of practice questions.

Ionic Nomenclature

Ions are atoms that have a charge because they have gained or lost electrons. **Ionic compounds** are formed when two ions of opposite charges are attracted to each other and form a bond. To name them, simply write the name of the cation (positive ion) followed by the name of the anion (negative ion).

Here is an example of a simple binary ionic compound. The anion in these compounds ends with *-ide*.



Calcium bromide

Polyatomic ions are made of multiple atoms. The anion will usually end with *-ate* or *-ite*.



Sodium nitrate

Any compound with a transition metal needs a Roman numeral to show the charge of the cation.



Zirconium (III) oxide

- | | | | |
|---------------------------------|--------------------|-----------------------------|------------------------------------|
| 1. MgS | Magnesium Sulfide | 4. NH_4Cl | Ammonium chloride |
| 2. KF | Potassium fluoride | 5. TiO_2 | Titanium dioxide |
| 3. $\text{Ca}_3(\text{PO}_4)_2$ | Calcium phosphate | 6. Cu_2CO_3 | carbon (II) carbonate
hydroxide |

Ionic Formulas

Ionic formulas must be written in such a way that the charges of the cation and anion balance and form a neutral compound. This can be done by adding **subscripts** to change the number of each ion within the compound.

⁺¹
Cesium ⁻³
nitride



Remember, Roman numerals after transition metals indicate the charge of the cation.

⁺²
Zinc (II) ⁻¹
hydroxide



- | | | | |
|------------------------|-------------------------|-------------------------|-----------------------------------|
| 7. Calcium chloride | CaCl_2 | 10. Iron (III) chlorite | FeCl_2 |
| 8. Aluminum phosphide | AlP | 11. Silver (I) oxalate | $\text{Ag}_2\text{C}_2\text{O}_4$ |
| 9. Potassium carbonate | K_2CO_3 | 12. Zinc (II) borate | $\text{B}_3\text{O}_6\text{Zn}_2$ |

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