

# Density Worksheet

## Physical Science

$D = m/V$

Densities of Common Substances @ 20°C				
Substance	Density (g/cm <sup>3</sup> )		Substance	Density (g/cm <sup>3</sup> )
Oxygen	0.00133		Aluminum	2.70
Hydrogen	0.000089		Iron	7.87
Ethanol	0.785		Copper	8.96
Benzene	0.880		Silver	10.5
Water	1.000		Lead	11.34
Magnesium	1.74		Mercury	13.6
Salt (sodium chloride)	2.16		Gold	19.32

1. The ratio of an object's mass to its \_\_\_\_\_ is called the density of the object.
2. A kilogram of lead occupies a much smaller volume than a kilogram of water, because \_\_\_\_\_ has a much higher density.
3. For the masses and volumes indicated, calculate the density in grams per cubic centimeters.
  - a. mass = 453 g; volume = 225 cm<sup>3</sup>
  - b. mass = 5.0 g; volume = 10.0 cm<sup>3</sup>
  - c. mass = 26.1 g; volume = 2.0 mL
4. If 89.2 mL of a liquid has a mass of 75.2 g, calculate the liquid's density.
5. A cube of metal weighs 1450 g and displaces 542 mL of water when immersed. Calculate the density of the metal.

# Physical Science If8767 Density

**RS Peters**



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