

**QUESTION 3**

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a  $0,2 \text{ mol.dm}^{-3}$  ethanoic acid solution to 10 g of  $\text{NaHCO}_3$  to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of  $\text{CO}_2$  produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in  $250 \text{ cm}^3$  of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3  $25 \text{ cm}^3$  of a potassium hydroxide solution of concentration  $0,25 \text{ mol.dm}^{-3}$  completely neutralises a dilute solution of sulphuric acid ( $\text{H}_2\text{SO}_4$ ) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

**[25]**

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**Thirumalaisamy P.**  
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