# Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- 1. The charge on an ion is
  - a. always positive.
  - b. always negative.
  - c. either positive or negative.
  - d. zero.
- According to the octet rule, a calcium atom has a tendency to
  - a. lose one electron.
  - b. lose two electrons.
  - c. gain one electron.
  - d. gain two electrons.
- 3. If a compound forms by ionic bonding, which is not true?
  - a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
  - The net charge of the compound is zero.
  - c. The compound contains just two atoms, each of opposite charge.
  - Several ions group together in a tightly packed structure.
- \_\_\_\_\_ 4. The only property listed that is not characteristic of ionic compounds is
  - a. high melting point.
  - b. hardness.
  - c. lack of crystal structure.
  - d. good conductor of electricity.
- 5. Which formula listed below represents a polyatomic ion?
  - a. NH<sub>4</sub>\*
  - b. H<sub>2</sub>SO<sub>4</sub>
  - c. NaCl.
  - d. H<sub>2</sub>O
- 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
  - a. ionic substances tend to vaporize at room temperature.
  - b. ionic substances are brittle.
  - attractive forces between ions are greater than the attractive forces between molecules.
  - d. the numbers of positive and negative charges are equal in an ionic compound.

# Physical Science Chemistry Test Chemical Bonds Answers

Siddappa N.Byrareddy

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## **Table of Contents Physical Science Chemistry Test Chemical Bonds Answers**

- 1. Understanding the eBook Physical Science Chemistry Test Chemical Bonds Answers
  - The Rise of Digital Reading Physical Science Chemistry Test Chemical Bonds Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Chemistry Test Chemical Bonds Answers
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Chemistry Test Chemical Bonds Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Chemistry Test Chemical Bonds Answers
  - Personalized Recommendations
  - Physical Science Chemistry Test Chemical Bonds Answers User Reviews and Ratings

- Physical Science Chemistry Test Chemical Bonds Answers and Bestseller Lists
- 5. Accessing Physical Science Chemistry Test Chemical Bonds Answers Free and Paid eBooks
  - Physical Science Chemistry Test Chemical Bonds Answers Public Domain eBooks
  - Physical Science Chemistry Test Chemical Bonds Answers eBook Subscription Services
  - Physical Science Chemistry Test Chemical Bonds Answers Budget-Friendly Options
- 6. Navigating Physical Science Chemistry Test Chemical Bonds Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Physical Science Chemistry Test Chemical Bonds Answers Compatibility with Devices
  - Physical Science Chemistry Test Chemical Bonds Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Chemistry Test Chemical Bonds Answers
  - Highlighting and Note-Taking Physical Science Chemistry Test Chemical Bonds Answers
  - Interactive Elements Physical Science Chemistry Test Chemical Bonds Answers
- 8. Staying Engaged with Physical Science Chemistry Test Chemical Bonds Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Chemistry Test Chemical Bonds Answers
- 9. Balancing eBooks and Physical Books Physical Science Chemistry Test Chemical Bonds Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Chemistry Test Chemical Bonds Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chemistry Test Chemical Bonds Answers
  - Setting Reading Goals Physical Science Chemistry Test Chemical Bonds Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chemistry Test Chemical Bonds Answers
  - Fact-Checking eBook Content of Physical Science Chemistry Test Chemical Bonds Answers
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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