

Chemistry MSL Review

Math in Chemistry

1. Give a definition and example of a qualitative measurement. Give a definition and example of a quantitative measurement.
2. How many sig figs are contained in the following numbers:
 0.067 2450 2450.0 0.0100
3. Define accuracy and precision.
4. If a block has a known density of 1.45, what is the percent error of the value 1.14 measured in lab?
5. A block of unknown composition has a mass of 23.2 grams. Using water displacement, its volume is measured at 27.9 mL. What is the block's density?
6. What is the volume of an object with a density of 1.94 g/mL and a mass of 15.5g?
7. Convert the temperature 88°C to Kelvin.
8. Metric conversions: Convert 4.58 m to cm. Convert 104.5 J to kJ. Convert 0.009 332 g to mg.
9. What is the difference between a chemical change and a physical change?

Atomic Structure

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.



1. Define the following terms:
 atomic number; mass number ; isotope ; atomic mass
2. How do you determine the number of protons in an atom? How do you determine the number of neutrons in an atom?
3. Complete the following table:

Isotope	Atomic #	# Protons	# Neutrons	Mass #
^{12}C				
	17		18	
		26		56
		17		37
			2	3
	52			128
		50	70	

- a. Consider the second and fourth row in the table above. What do they have in common?
- b. Consider the second and fourth row in the table above. What are the differences?

North Carolina Chemistry Msl Test Study Guide

Karin Nielsen-Saines



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Table of Contents North Carolina Chemistry Msl Test Study Guide

1. Understanding the eBook North Carolina Chemistry Msl Test Study Guide
 - The Rise of Digital Reading North Carolina Chemistry Msl Test Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying North Carolina Chemistry Msl Test Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an North Carolina Chemistry Msl Test Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from North Carolina Chemistry Msl Test Study Guide
 - Personalized Recommendations
 - North Carolina Chemistry Msl Test Study Guide User Reviews and Ratings
 - North Carolina Chemistry Msl Test Study Guide and Bestseller Lists
5. Accessing North Carolina Chemistry Msl Test Study Guide Free and Paid eBooks

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 - Interactive Elements North Carolina Chemistry Msl Test Study Guide
8. Staying Engaged with North Carolina Chemistry Msl Test Study Guide
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers North Carolina Chemistry Msl Test Study Guide
9. Balancing eBooks and Physical Books North Carolina Chemistry Msl Test Study Guide
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection North Carolina Chemistry Msl Test Study Guide
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine North Carolina Chemistry Msl Test Study Guide
- Setting Reading Goals North Carolina Chemistry Msl Test Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of North Carolina Chemistry Msl Test Study Guide
- Fact-Checking eBook Content of North Carolina Chemistry Msl Test Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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