

**GR 10 Paper 1 Life Sciences scope November 2023 (10 November 2023)**

<b>Biological terms:</b> Interphase, Anaphase, Metaphase, Telophase, Collenchyma, phloem, Sclerenchyma, Parenchyma, cuticula, Transpirations, stomata, neuron, organic compound, connective tissue, chromatin, chromosome, foramen magnum, axial skeleton, cranium	
<b>Chemistry of life</b>	<b>Cells</b>
Difference between organic and inorganic compounds	Difference between animal cell and plant cell
Examples of organic and inorganic compounds	All the cell organelles and functions/rolls
Monomers of all organic compound	Describe the build of the nucleus, mitochondria, and chloroplast
Elements of all the organic compounds	<b>Investigation:</b> Osmose and diffusion (Look at the apparatus that we use and how to set up this Investigation)
Vitamins A, B, C, D, and E, and there deficiency diseases they can cause	Explain osmose and diffusion
<b>Enzymes</b>	
- Lock and key model	
- Function of enzymes	
Influence of temperature and pH on enzyme action	
<b>Mitosis</b>	<b>Plant tissue</b>
All the phases can identify and give reasons for your answers	<b>Build and function of:</b> Collenchyma, Sclerenchyma, Parenchyma, Phloem, Xylem and stomata
Product of Mitosis	Definition of tissue
Difference between telophase in plant cells and animal cells	
Biological importance of Mitosis	
<b>Cancer: - Uncontrolled cell division and growth:</b> Causes and treatment of cancer	
	<b>Function and build of:</b> Stomata, Xylem, and root hair
	<b>Adaptions of Xylem for movement of water upwards</b>
<b>Animal tissue</b>	<b>Plant organs</b>
Name and identify all the animal tissues and where in the human body are they located.	The labels and functions of the dorsiventral leaf.
Functions of all the animal tissues	Adaptions of the different parts of the leaf
Name all the connective tissues	Definition of an organ
<b>Support and transport in plants</b>	<b>Support in animals</b>
Labels of dicotyledonous stem and root	All the labels of the skeleton
Difference between a dicotyledons root and stem	Where do we find long bones in the skeleton
Investigation: The rate of transpirations (Potometer)	Function of the skeleton
The effect of environmental factors on transpiration (Wind, Temperature and Light.)	

# Scope For Life Sciences Paper 1 2014 Final Examination

**D Siedentop**



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web google classroom for the following redox equation  $\text{MnO}_4^- + 4\text{H}^+ + \text{aq} + \text{SO}_2 + 2\text{H}_2\text{O} + \text{aq} + \text{H}_2\text{SO}_4 + 4\text{H}^+ + \text{aq}$  what will be the stoichiometric coefficient of  $\text{H}_2\text{SO}_4$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}^{2+} + \text{Fe}^{3+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{2+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow 2\text{CNO} + \text{Au}^{3+} + 2\text{e}^-$   $2\text{H}_2\text{O}_2 + 2\text{e}^- + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O}$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{H}_2\text{O} \rightarrow 4\text{CNO} + 2\text{Au}^{3+} + 2\text{OH}^-$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 + 2\text{OH}^- \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh ions or the oh h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo and cr oh<sub>4</sub> in basic solution created by jay

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web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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redox equations oxidation state method ion electron method the latter is easier to use with redox reactions in aqueous solution and if necessary can be adapted to many situations that are not in aqueous solution

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add  $\text{H}^+$  ions or the  $\text{H}_2\text{O}$  pair to fully balance the equation in this video we'll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride  $\text{Cl}^-$  ions in acidic solution created by jay

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_2 \rightarrow \text{Cr}_2\text{O}_4^{2-} + \text{MnO}_2$   $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_2 \rightarrow \text{Cr}_2\text{O}_4^{2-} + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} \rightarrow \text{TeO}_3 + \text{N}_2\text{O}_4$  9  $\text{O}_2 + \text{Cr}_3\text{Te} \rightarrow \text{TeO}_4$  10  $\text{PbO}_2 + \text{IO}_2 \rightarrow \text{Pb}_2\text{O}_3$  11  $\text{PbO}_2 + \text{IO}_2 \rightarrow \text{Pb}_2\text{O}_3$  12  $\text{AsO}_3 + \text{H}_2\text{AsO}_4$

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding  $\text{H}_2\text{O}$  and  $\text{H}^+$  in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$  choose 1 answer c  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  b  $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  h and  $\text{NO}_2$  c h and

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