

16 • Chemical Equilibrium

PRACTICE TEST

1. Consider the reaction system,
 $\text{CoO}(s) + \text{H}_2(g) \rightleftharpoons \text{Co}(s) + \text{H}_2\text{O}(g)$ [Products] [Reactants]

The equilibrium constant expression is

- a) $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$ d) $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$ [H₂O]
 b) $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$ [H₂]
 c) $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{H}_2]}$ [H₂]

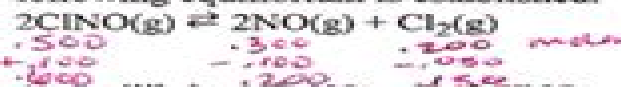
2. Given the equilibrium,
 $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, if this as soon as
 equilibrium is established by beginning with even 1 mol
 equal number of moles of SO_2 and O_2 in a occurs [SO₂]
 1.0 Liter bulb, then the following < [O₂] must be

true at equilibrium:

- a) $[\text{SO}_2] = [\text{SO}_3]$ (d) $[\text{SO}_2] < [\text{O}_2]$
 b) $2[\text{SO}_2] = 2[\text{SO}_3]$ c) $[\text{SO}_2] > [\text{O}_2]$
 c) $[\text{SO}_2] = [\text{O}_2]$

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole NO,
 0.200 mol Cl_2 and 0.500 mol ClNO were
 placed in a 25.0 Liter container. The
 following equilibrium is established:



3. At equilibrium, 0.600 mol of ClNO was
 present. The number of moles of Cl_2 present
 at equilibrium is

- a) 0.050 d) 0.200
 b) 0.100 e) 0.250
(c) 0.150

4. The equilibrium constant, K_c , is:

- a) 4.45×10^{-4} d) 0.167 = $\left(\frac{.200}{25}\right)^2 \left(\frac{.150}{25}\right)$
(b) 6.67×10^{-4} e) 1500 = $\frac{(.600/25)^2}{(.150/25)}$
 c) 0.111

5. At 985°C, the equilibrium constant for the
 reaction,



is 1.63. What is the equilibrium constant for
 the reverse reaction? K' = $\frac{1}{K} = \frac{1}{1.63}$

- a) 1.63 (d) 0.613 = $\frac{1}{1.63}$
 b) 0.815 e) 1.00
 c) 2.66

6. What is the relationship between K_p and K_c
 for the reaction, $2\text{ICl}(g) \rightleftharpoons \text{I}_2(g) + \text{Cl}_2(g)$?

- a) $K_p = K_c(RT)^{-1}$ (d) $K_p = K_c$ Δn = 0
 b) $K_p = K_c(RT)$ e) $K_p = K_c(2RT)$
 c) $K_p = K_c(RT)^2$

7. For the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$, K_p at
 25°C is 7.3, when all partial pressures are
 expressed in atmospheres. What is K_c for
 this reaction? Δn = 1 - 2 = -1
298 K [R = 0.0821 L·atm·mol⁻¹·K⁻¹]

- a) 4270 (d) 179 K_p = K_c(RT)^{Δn}
 b) 0.0119 e) 2.06 K_c = $\frac{K_p}{(RT)^{Δn}}$
 c) 0.291

8. 0.200 mol NO is placed in a one liter flask at
 2273 K. After equilibrium is attained, 0.0863
 mol N_2 and 0.0863 mol O_2 are present. What
 is K_c for this reaction?



- (a) 9.92 d) 39.7
 b) 3.15 e) 0.576
 c) 0.0372

.200 M	0	0
-2(.0863)	-.0863	-.0863
.0274	.0863	.0863

$K_c = \frac{(.0863)(.0863)}{(.0274)^2} = 9.92$

Review Chemical Equilibrium Answers

Guillaume Favre



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