

Unit 8 Review

Write *increases* or *decreases* to complete each relationship.

1. As volume increases, pressure _____decreases_____.
2. As volume increases, temperature _____increases_____.
3. As pressure increases, temperature _____increases_____.
4. As pressure increases, concentration _____increases_____.
5. As force decreases, pressure _____decreases_____.

True or False.

6. T The Kinetic Theory of Matter deals with moving particles.
7. T Gases can be collected by trapping them over water.
8. T Real gases exist in the world around us.
9. T Densities of gases are lower than 1 g/mL.
10. F If the barometric pressure is falling, that means good weather is coming.
11. T The troposphere contains the most air particles.
12. T Temperature is a measure of how fast particles are moving.

13. The gas that makes up the highest percentage of the atmosphere is

 nitrogen .

14. Atmospheric pressure is measured using a barometer .

15. Squeezing gas particles together is known as compressibility .

16. When particles collide with the sides of a container, they create pressure .

17. The spreading of a gas through a small opening is effusion .

18. The number of gas particles present is called the concentration .

Gas Law Problems. Write the answer with the correct unit.

19. A gas has a beginning pressure of 2.95 atm and a volume of 5.15 L. If the ending pressure is 4.23 atm, what is the ending volume? $P_1V_1 = P_2V_2$
 $(2.95)(5.15) \div 4.23 = 3.59$

20. A gas has a beginning temperature of 300 K and a volume of 4.06 L. If the ending temperature is 176 K, what is the ending volume? $V_1T_2 = V_2T_1$
 $(4.06)(176) \div 300 = 2.38$

21. A gas has a beginning pressure of 2.09 atm, a temperature of 268 K, and a volume of 0.450 L. If the ending pressure is 1.2 atm and the temperature is 149 K, what is the ending volume? $P_1V_1T_2 = P_2V_2T_1$

Review Section 1 Gases Answers

Cynthia Y. Young



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