

Exam 2 Review (Chapters 12 and 13)

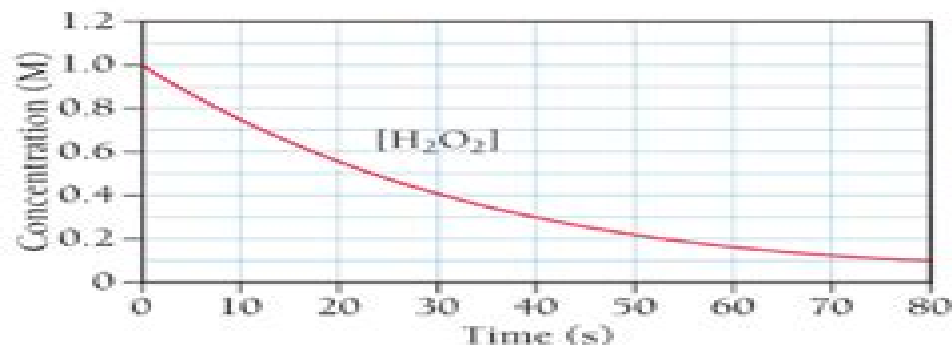
(Disclaimer: this review is made up of questions pulled from previous SI Leaders reviews and should not be your only study source for the upcoming exam. It is just a way to test your knowledge of what I thought was important. Also, just like you I can make mistakes, so if you find any please let me know!)

1. Consider the balanced equation below



Write the rate of the following reaction in terms of the appearance of products and disappearance of reactants.

2.



The graph above shows the concentration of H_2O_2 as a function of time for the reaction $2\text{H}_2\text{O}_2 (\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2 (\text{g})$. Using the graph complete the Following.

- a) Find the average rate of the reaction between 10 and 20 seconds

Review Module Chapters 12 Answer

Scott Selikoff, Jeanne Boyarsky



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