

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 11 REVIEW***Gases***SECTION 3****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ The molar mass of a gas at STP is the density of that gas  
(a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L.  
(b) divided by the mass of 1 mol. (d) divided by 22.4 L.
2. \_\_\_\_\_ For the expression  $V = \frac{nRT}{P}$ , which of the following will cause the volume to increase?  
(a) increasing  $P$  (c) increasing  $T$   
(b) decreasing  $T$  (d) decreasing  $n$
3. Two sealed flasks, A and B, contain two different gases of equal volume at the same temperature and pressure.  
\_\_\_\_\_ a. The two flasks must contain an equal number of molecules. True or False?  
\_\_\_\_\_ b. The two samples must have equal masses. True or False?

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

4. Use the data in the table below to answer the following questions.

| Formula                       | Molar mass (g/mol) |
|-------------------------------|--------------------|
| N <sub>2</sub>                | 28.02              |
| CO                            | 28.01              |
| C <sub>2</sub> H <sub>2</sub> | 26.04              |
| He                            | 4.00               |
| Ar                            | 39.95              |

(Assume all gases are at STP.)

- \_\_\_\_\_ a. Which gas contains the most molecules in a 5.0 L sample?  
\_\_\_\_\_ b. Which gas is the least dense?  
\_\_\_\_\_ c. Which two gases have virtually the same density?  
\_\_\_\_\_ d. What is the density of N<sub>2</sub> measured at STP?

# Review Gases Section Short Answer

**Liying Dong**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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