

**13.1****THE NATURE OF GASES****Section Review****Objectives**

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

**Vocabulary**

- kinetic energy
- kinetic theory
- gas pressure
- vacuum
- atmospheric pressure
- barometer
- pascal (Pa)
- standard atmosphere (atm)

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The kinetic theory describes the \_\_\_\_\_ of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly \_\_\_\_\_, that the particles of gas are relatively \_\_\_\_\_, move \_\_\_\_\_ of each other, and are in constant \_\_\_\_\_ motion. The \_\_\_\_\_ between particles are perfectly elastic so that the total \_\_\_\_\_ remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure \_\_\_\_\_ pressure. Standard conditions are defined as a temperature of \_\_\_\_\_ and a pressure of \_\_\_\_\_.

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**Part B True-False**

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- \_\_\_\_\_ 11. Atmospheric pressure is 760 mm Hg.
- \_\_\_\_\_ 12. The SI unit of pressure is the pascal.

# Review Gases Section 1 Answers

**Victor M. Corman**



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