

AP® CHEMISTRY 2005 EXAMINATIONS CHEMISTRY SECTION

Question 10
(8 points)

As sample of $\text{C}_2\text{H}_2\text{Cl}_2$ is subjected to electrolysis, it is found that 0.010 moles of chlorine gas are formed for every 0.0050 moles of the sample. An electrochemical cell is constructed as follows:



Use the standard E° in brackets for the reduction half-reaction. The temperature of the $\text{C}_2\text{H}_2\text{Cl}_2$ is 25°C. (Assume that the activity of the ions is unity.)

Answer each part:

- Write the half-reactions occurring. Show enough detail to indicate the balanced redox reaction. Indicate whether the reaction is spontaneous or nonspontaneous.
- Write the half-reactions occurring. Show enough detail to indicate the balanced redox reaction. Indicate whether the reaction is spontaneous or nonspontaneous.

1 point is earned for each half-reaction.

$\text{C}_2\text{H}_2\text{Cl}_2$ reacts readily with H_2O to produce $\text{C}_2\text{H}_4\text{Cl}_2$, as represented by the following equation:



The $\text{C}_2\text{H}_2\text{Cl}_2$ is a liquid at room temperature, whereas $\text{C}_2\text{H}_4\text{Cl}_2$ is a liquid at room temperature. Ethane is the only hydrocarbon to boil at a higher temperature than the hydrocarbon formed in the reaction.

- Write the Lewis structure of the product, $\text{C}_2\text{H}_4\text{Cl}_2$. Indicate all electron pairs and the hybridization of the carbon atoms.
- Write the Lewis structure of the product, $\text{C}_2\text{H}_4\text{Cl}_2$. Indicate all electron pairs and the hybridization of the carbon atoms.

1 point is earned for each structure.

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 - The Rise of Digital Reading Response Question Ap Chemistry 2013
 - Advantages of eBooks Over Traditional Books
2. Identifying Response Question Ap Chemistry 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Response Question Ap Chemistry 2013
 - User-Friendly Interface
4. Exploring eBook Recommendations from Response Question Ap Chemistry 2013
 - Personalized Recommendations
 - Response Question Ap Chemistry 2013 User Reviews and Ratings
 - Response Question Ap Chemistry 2013 and Bestseller Lists

5. Accessing Response Question Ap Chemistry 2013 Free and Paid eBooks
 - Response Question Ap Chemistry 2013 Public Domain eBooks
 - Response Question Ap Chemistry 2013 eBook Subscription Services
 - Response Question Ap Chemistry 2013 Budget-Friendly Options
6. Navigating Response Question Ap Chemistry 2013 eBook Formats
 - ePub, PDF, MOBI, and More
 - Response Question Ap Chemistry 2013 Compatibility with Devices
 - Response Question Ap Chemistry 2013 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Response Question Ap Chemistry 2013
 - Highlighting and Note-Taking Response Question Ap Chemistry 2013
 - Interactive Elements Response Question Ap Chemistry 2013
8. Staying Engaged with Response Question Ap Chemistry 2013
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Response Question Ap Chemistry 2013
9. Balancing eBooks and Physical Books Response Question Ap Chemistry 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Response Question Ap Chemistry 2013
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Response Question Ap Chemistry 2013
 - Setting Reading Goals Response Question Ap Chemistry 2013
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Response Question Ap Chemistry 2013
 - Fact-Checking eBook Content of Response Question Ap Chemistry 2013
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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