

Question	Answer	Marks	Guidance
5 (a)	F_2 forces mark F_2 has van der Waals' (forces) OR F_2 has induced dipole attractions OR interactions OR F_2 has temporary OR instantaneous dipole(-dipole) attraction OR interactions ✓ HCl forces mark HCl has permanent dipole(-dipole) attractions OR interactions ✓ Comparison of strength of forces between molecules mark intermolecular force in HCl is stronger than that in F_2 OR permanent dipoles are stronger (than induced dipoles) ✓ Boiling point mark more energy is required to break stronger (intermolecular) forces ✓	4	Use annotations ie ticks crosses ECF ^ etc for this part ALLOW vdWs for van der Waals' IGNORE F_2 has covalent bond for this mark IGNORE F_2 has 'intermolecular forces' Quality of written communication: 'dipole(s)' spelled correctly and used in context for the second marking point IGNORE HCl has 'intermolecular forces' IGNORE van der Waals' forces in HCl/ DO NOT ALLOW hydrogen bonding DO NOT ALLOW ionic bonding Look for strength of force comparison anywhere in the answer ALLOW ECF for hydrogen bonding in HCl being stronger than the stated intermolecular forces in F_2 BUT DO NOT ALLOW this mark if HCl or F_2 has covalent bonds broken OR if HCl has ionic bonds broken (the question asks for forces between molecules) IGNORE HCl has stronger van der Waals' (forces) than F_2 (as they both have the same number of electrons) DO NOT ALLOW fourth mark if covalent bonds are broken in HCl or F_2 OR if ionic bonds are broken in HCl IGNORE 'heat' but ALLOW 'heat energy'

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