

7. Chromite is the main ore of chromium. The chromium-containing compound in chromite is FeCr_2O_4 . The percentage of chromium in a sample of chromite can be determined using the method below.

Step 1

A 5.28 g sample of chromite ore is heated with sodium peroxide, Na_2O_2 .



Water is added to the resulting mixture.

Na_2O_2 dissolves in the water forming a solution containing CO_3^{2-} ions.

Step 2

The mixture from Step 1 is filtered and the filtrate is made up to 1.00 dm³ in a volumetric flask.

A 25.0 cm³ sample of this alkaline solution is pipetted into a conical flask and an excess of aqueous potassium iodide is added.

- A redox reaction takes place between I^- ions, CO_3^{2-} ions and H_2O .
- In this reaction 1 mol CO_3^{2-} reacts 1.5 mol I_2 .

Step 3

The resulting mixture is titrated with 0.100 mol dm⁻³ sodium thiosulfate, $\text{Na}_2\text{S}_2\text{O}_3$ (aq) to estimate the I_2 present.



The average titre of $\text{Na}_2\text{S}_2\text{O}_3$ (aq) is 26.5 cm³.

- (a) In Step 1, Na_2O_2 and NaFeO_2 react with water forming an alkaline solution containing a brown precipitate. This is not a redox reaction.

Write equations for:

- the reaction of Na_2O_2 with water
- the reaction of NaFeO_2 with water.



$$\begin{array}{rcl} \text{I}^- + 2 + (2 \times 2) & = & 0 \\ \text{I}^- & - & 1 \\ \hline & & -3 \end{array}$$

- (b) Calculate the percentage, by mass, of chromium in the ore.

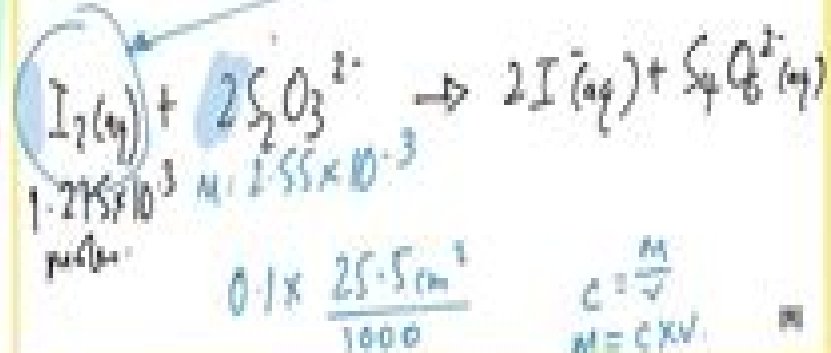
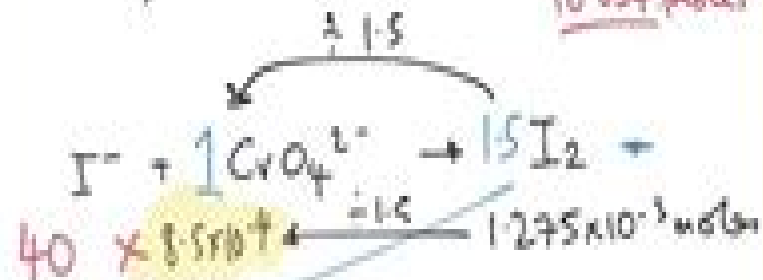
Give your answer to one decimal place.

$$\text{mass \%} = \frac{\text{mass of Cr}}{\text{mass of ore}} \times 100$$

$$\text{mass \%} = \frac{10.034 \text{ g}}{5.28 \text{ g}} \times 100 = 52.0$$



10.034 moles



- (c) This part refers to Step 2 of the method.

Write the ionic equation between I^- ions, CO_3^{2-} ions and H_2O .

- CO_3^{2-} ions are reduced to carbonate ions.
- I^- ions are oxidised to make I_2 .
- Construct an overall equation for the redox reaction and write half equations for the oxidation and reduction.

Overall equation:

Half equations:

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