

Part 2

GCSE Chemistry OCR A (Combined Science - **HIGHER**) Paper 2 Exam Walkthrough

Combined Science Higher - Chemistry

Walkthrough Mark Paper

12
19

Look at the equation for the equilibrium reaction between sulfur dioxide, SO_2 , oxygen and sulfur trioxide, SO_3 , in a closed system.



- (a) Predict the effect of adding more oxygen to the equilibrium mixture.

Explain your answer.

(2)

- (b) Predict the effect of increasing the total pressure in the equilibrium mixture.

Explain your answer.

(2)

- (c) Predict the effect of increasing the **temperature** of the equilibrium mixture.

Explain your answer.

(2)

- (d) Sulfur dioxide for this reaction is made by burning sulfur.



Calculate the mass of sulfur needed to make 48 tonnes of sulfur dioxide.

Give your answer to 2 significant figures.

The relative atomic mass, A_r , of S is 32.1 and of O is 16.0.

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JG Myers



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put the loaf tray on the middle shelf and bake for 25 minutes then lower the oven to 200 c 392 f gas mark 6 and bake for 10 15 minutes more until the crust has a good colour hold the loaf in a tea towel and tap the bottom if it sounds hollow it s

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meanwhile heat the oven to 230 c 210 c fan 450 f gas 8 and put a roasting tray in the bottom of the oven to heat up step 9

remove the risen loaf from the bag and dust with flour fill the roasting tray with cold water to create steam and put the bread in the oven bake for 15 minutes then lower the oven to 190 c 170 c fan 375 f gas 5

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