

- (b) A solution is prepared by the addition of 10.0 mL of 0.100 M CH_3COOH and 10.0 mL of 0.100 M NaOH .

- (c) The solution is a buffer solution of CH_3COOH and CH_3COO^- .

The total volume is 20.0 mL. The concentration of CH_3COOH is 0.050 M.

In this case, we should consider both of the species involved that could be acting as a weak acid or base.

CH_3COOH is a weak acid with $K_a = 1.8 \times 10^{-5}$ and CH_3COO^- is a weak base.

$$[\text{CH}_3\text{COOH}] = 0.050 \text{ M} \quad [\text{CH}_3\text{COO}^-] = 0.050 \text{ M}$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad [\text{CH}_3\text{COO}^-] = \frac{[\text{CH}_3\text{COOH}][\text{H}^+]}{K_a} = \frac{0.050 \text{ M} [\text{H}^+]}{1.8 \times 10^{-5}}$$

$$\text{pH} = -\log [\text{H}^+] = 4.74 \quad \text{pH} = 4.74$$

$$\text{pH} = 4.74 \quad \text{pH} = 4.74$$

Reaction: $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$
Initial: 0.100 M, 0.100 M, 0.000 M, 0.000 M
Change: -0.050 M, -0.050 M, +0.050 M, +0.050 M
Equilibrium: 0.050 M, 0.050 M, 0.050 M, 0.050 M
pH = 4.74

- (b) The solution is a buffer solution of CH_3COOH and CH_3COO^- . The concentration of CH_3COOH is 0.050 M and the concentration of CH_3COO^- is 0.050 M.

In this case, we should consider both of the species involved that could be acting as a weak acid or base.



- (c) The solution is a buffer solution of CH_3COOH and CH_3COO^- . The concentration of CH_3COOH is 0.050 M and the concentration of CH_3COO^- is 0.050 M.

In this case, we should consider both of the species involved that could be acting as a weak acid or base.

- (d) The solution is a buffer solution of CH_3COOH and CH_3COO^- .

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad [\text{CH}_3\text{COO}^-] = \frac{[\text{CH}_3\text{COOH}][\text{H}^+]}{K_a} = \frac{0.050 \text{ M} [\text{H}^+]}{1.8 \times 10^{-5}}$$

$$\text{pH} = -\log [\text{H}^+] = 4.74 \quad \text{pH} = 4.74$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad [\text{CH}_3\text{COO}^-] = \frac{[\text{CH}_3\text{COOH}][\text{H}^+]}{K_a} = \frac{0.050 \text{ M} [\text{H}^+]}{1.8 \times 10^{-5}}$$

$$\text{pH} = 4.74$$

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Mike Smith



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