

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the state of dynamic equilibrium involving interconverting reactants simultaneously forming as reactants or products and reactants. These reactions are called reversible reactions, as the reaction can proceed in either direction. Once the equilibrium is reached the forward and reverse reactions continue at the same rate and the concentrations of reactants and products remain constant. The importance of understanding that equilibrium is a dynamic state is that it allows us to predict the behavior of chemical systems. For example, if a reaction is at equilibrium, then the forward and reverse reactions continue to occur over time, but there is no net change in the concentrations of reactants and products. Equilibrium conditions are reached only at constant pressure, or concentration, or volume. Le Chatelier's Principle states that applying stress changes the equilibrium position of a system in order to minimize the effect of the change and bring the system back to equilibrium. When more products are added, an equilibrium reaction will shift to the left to consume the added products and bring the system back to equilibrium. When more reactants are added, the equilibrium will shift to the right to consume the added reactants and bring the system back to equilibrium. In this lab, you will be observing the changes in equilibrium when the amount of reactants, and temperature are changed. Follow Le Chatelier's Principle in the following reactions. Shifts in equilibrium are indicated by the color changes in solution according to the color indicators in the equations below. For example, if the color changes to red, the reaction shifts to the right to consume the added reactants and bring the system back to equilibrium.



Adding NaOH solution to $\text{Fe}^{3+}(\text{aq})$ and $\text{SCN}^{-}(\text{aq})$ will allow Fe^{3+} ions to react with OH^{-} ions, leaving the SCN^{-} ions in solution. This is an example of a shift to the right to consume the added reactants and bring the system back to equilibrium.

Procedure (This Group):

- | | |
|--|--|
| 1. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 2. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 3. Prepare 100 mL of 0.1 M NaOH solution. | 4. Prepare 100 mL of 0.1 M HCl solution. |
| 5. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 6. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 7. Prepare 100 mL of 0.1 M NaOH solution. | 8. Prepare 100 mL of 0.1 M HCl solution. |
| 9. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 10. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 11. Prepare 100 mL of 0.1 M NaOH solution. | 12. Prepare 100 mL of 0.1 M HCl solution. |
| 13. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 14. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 15. Prepare 100 mL of 0.1 M NaOH solution. | 16. Prepare 100 mL of 0.1 M HCl solution. |

Observing Chemical Equilibrium Lab 46

**William Crowell Bray, Wendell Mitchell
Latimer**



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Chemistry Lab Manual Class XI | follows the latest CBSE syllabus and other State Board following the CBSE Curriculum. Mr. Rohit Manglik, Mr. Kaushalesh Dwivedi, 2022-08-04 With the NEP and expansion of research and knowledge has changed the face of education to a great extent In the Modern times education is not just constricted to the lecture method but also includes a practical knowledge of certain subjects This way of education helps a student to grasp the basic concepts and principles Thus trying to break the stereotype that subjects like Physics Chemistry and Biology means studying lengthy formulas complex structures and handling complicated instruments we are trying to make education easy fun and enjoyable

Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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articles to detailed theoretical papers provide rigorous supporting material so that the reader can delve as deeply as desired into the subject International Aerospace Abstracts ,1999 Experiment Station Record United States. Office of Experiment Stations,1913 **The Iron(III) Thiocyanate Reaction** Kevin C. de Berg,2019-11-13 This Brief presents an historical investigation into the reaction between ferric ions and thiocyanate ions which has been viewed in different ways throughout the last two centuries Historically the reaction was used in chemical analysis and to highlight the nature of chemical reactions the laws of chemistry models and theories of chemistry chemical nomenclature mathematics and data analysis and instrumentation which are important ingredients of what one might call the nature of chemistry Using the history of the iron III thiocyanate reaction as a basis the book s main objective is to explore how chemistry develops its own knowledge base how it assesses the reliability of that base and how some important tools of the trade have been brought to bear on a chemical reaction to achieve understanding a worthwhile goal of any historical investigation **Essential Practical NMR for Organic Chemistry** S. A. Richards,J. C. Hollerton,2023-04-03 Essential Practical NMR for Organic Chemistry A hands on resource advocating an ordered approach to gathering and interpreting NMR data The second edition of Essential Practical NMR for Organic Chemistry delivers a pragmatic and accessible text demonstrating an ordered approach to gathering and interpreting NMR data In this informal guide you ll learn to make sense of the high density of NMR information through the authors problem solving strategies and interpretations The book also discusses critical aspects of NMR theory as well as data acquisition and processing strategy It explains the use of NMR spectroscopy for dealing with problems of small organic molecule structural elucidation and includes a brand new chapter on Nitrogen 15 NMR Readers will also find Strategies for preparing a sample spectrum acquisition processing and interpreting your spectrum Fulsome discussions of Carbon 13 NMR spectroscopy Practical treatments of quantification safety procedures and relevant software An ideal handbook for anyone involved in using NMR to solve structural problems this latest edition of Essential Practical NMR for Organic Chemistry will be particularly useful for chemists running and looking at their own NMR spectra as well as those who work in small molecule NMR It will also earn a place in the libraries of undergraduate and post graduate organic chemistry students Modern Nmr Techniques and Their Application in Chemistry Millicent Popov,1990-09-28 Details theory methods and successful application of NMR technology for chemists chemical engineers and biochemists those with specialized knowledge of NMR spectroscopy and those who want to catch up with the spectacular evolution of the field during the last decade Reviews one dimensional NMR *Observation, Prediction and Simulation of Phase Transitions in Complex Fluids* Marc Baus,L.F Rull,Jean-Paul Ryckaert,2012-12-06 Observation Prediction and Simulation of Phase Transitions in Complex Fluids presents an overview of the phase transitions that occur in a variety of soft matter systems colloidal suspensions of spherical or rod like particles and their mixtures directed polymers and polymer blends colloid polymer mixtures and liquid forming mesogens This modern and fascinating branch of condensed matter physics is presented

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