

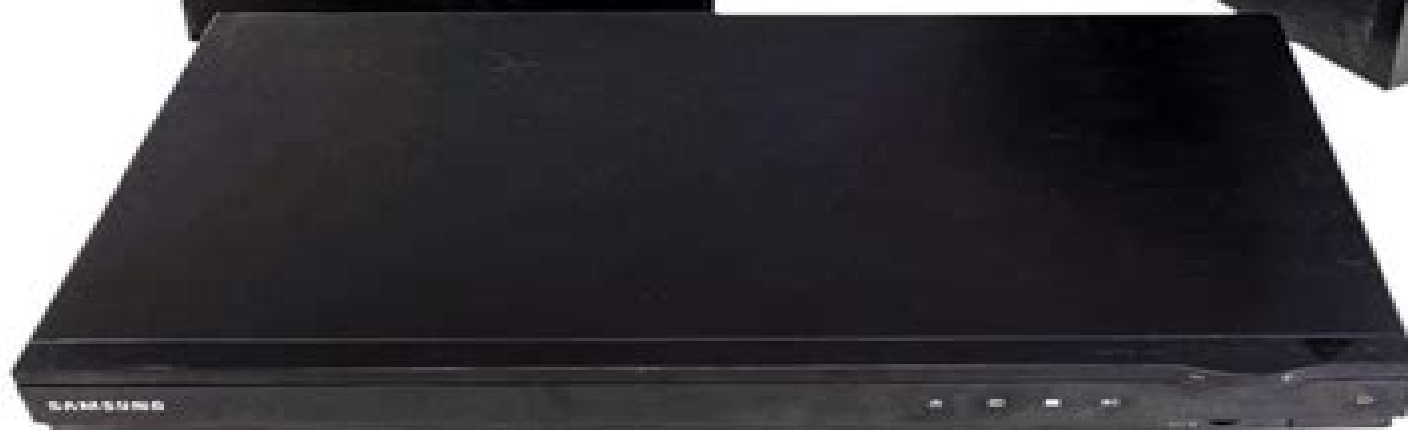
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$$\begin{array}{l} 1 \text{ n}_2 + \text{ f}_2 \rightarrow 2 \text{ nfs} \\ 2 \text{ c}_6\text{h}_{10} + \text{ o}_2 \rightarrow \text{ co}_2 + \text{ h}_2\text{O} \\ 3 \text{ hbr} + \text{ khco}_3 + \text{ h}_2\text{O} \rightarrow \text{ kbr} + \text{ co}_2 + \text{ gabr} + \text{ nabr} + \text{ sno}_4 + \text{ naso}_3 + \text{ gaz} + \text{ so}_3 \\ 3 \text{ 5 nf}_3 + \text{ snf}_2 + \text{ n}_2\text{O}_3 \text{ using the equation from problem 2 above answer the following} \end{array}$$

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$$\begin{array}{l} 1 \text{ 1 n}_2 + 3 \text{ f}_2 \rightarrow 2 \text{ nf}_3 \\ 2 \text{ 2 c}_6\text{h}_{10} + 17 \text{ o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{O} \\ 3 \text{ 4 5 1 hbr} + 1 \text{ khco}_3 + 1 \text{ h}_2\text{O} + 1 \text{ kbr} + 2 \text{ co}_2 + 2 \text{ gabr} + 3 \text{ 3 na}_2\text{so}_3 + 1 \text{ ga}_2\text{so}_3 + 3 \text{ 6 nabr} + 3 \text{ sno}_2 + \text{ nf}_3 + 3 \text{ snf}_2 + 1 \text{ n}_2\text{O}_3 \text{ using the following equation} \\ 2 \text{ naoh} + \text{ h}_2\text{so}_4 \rightarrow 2 \text{ h}_2\text{O} + \text{ na}_2\text{so}_4 \end{array}$$

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assuming CeCl_2 CuCl_2 is available in excess

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{H}_2\text{Br}_2$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{O} + \text{Na}_2\text{SO}_3$
 $\text{Ga}_2\text{O}_3 + \text{NaBr} \rightarrow \text{Na}_3\text{GaO}_4 + \text{Br}_2$
 using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{Ga}_2\text{O}_3 + \text{NaBr} \rightarrow \text{Na}_3\text{GaO}_4 + \text{Br}_2$
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 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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