

Three-mechanism reactions

10. A reaction $\text{A} \rightarrow \text{B}$ has $E_a = 100 \text{ kJ mol}^{-1}$. At 25°C , $k = 1.0 \times 10^{-5} \text{ s}^{-1}$. At 35°C , $k = 1.0 \times 10^{-4} \text{ s}^{-1}$.

Which statement is correct for this reaction mechanism?

- a. A and B are tetrahedral species of MgCl_2 type
- b. A , B and $\text{B}^\ddagger$ are tetrahedral species of MgCl_2 type
- c. A and B are tetrahedral species of MgCl_2 type
- d. A and B are tetrahedral species of MgCl_2 type
- e. A and B are tetrahedral species of MgCl_2 type

11. The energy profile diagram for the reaction $\text{A} \rightarrow \text{B}$ is shown below. The energy of A is 100 kJ mol^{-1} .

What is the energy of B ?

12. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

What is the half-life of A ?

13. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

14. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

15. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

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17. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

- a. $1.0 \times 10^{-5} \text{ s}^{-1}$
- b. $1.0 \times 10^{-4} \text{ s}^{-1}$
- c. $1.0 \times 10^{-3} \text{ s}^{-1}$
- d. $1.0 \times 10^{-2} \text{ s}^{-1}$

18. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

19. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

20. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

- a. $1.0 \times 10^{-5} \text{ s}^{-1}$
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21. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

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24. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

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26. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

- a. $1.0 \times 10^{-5} \text{ s}^{-1}$
- b. $1.0 \times 10^{-4} \text{ s}^{-1}$
- c. $1.0 \times 10^{-3} \text{ s}^{-1}$
- d. $1.0 \times 10^{-2} \text{ s}^{-1}$

27. The reaction $\text{A} \rightarrow \text{B}$ is first order. The rate constant is $1.0 \times 10^{-5} \text{ s}^{-1}$.

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1. Understanding the eBook Review Questions Thermochem
 - The Rise of Digital Reading Review Questions Thermochem
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Questions Thermochem
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Questions Thermochem
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Questions Thermochem
 - Personalized Recommendations
 - Review Questions Thermochem User Reviews and Ratings
 - Review Questions Thermochem and Bestseller Lists
5. Accessing Review Questions Thermochem Free and Paid eBooks
 - Review Questions Thermochem Public Domain eBooks
 - Review Questions Thermochem eBook Subscription Services
 - Review Questions Thermochem Budget-Friendly Options
6. Navigating Review Questions Thermochem eBook Formats

- ePub, PDF, MOBI, and More
- Review Questions Thermochem Compatibility with Devices
- Review Questions Thermochem Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Questions Thermochem
 - Highlighting and Note-Taking Review Questions Thermochem
 - Interactive Elements Review Questions Thermochem
- 8. Staying Engaged with Review Questions Thermochem
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Questions Thermochem
- 9. Balancing eBooks and Physical Books Review Questions Thermochem
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Questions Thermochem
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Questions Thermochem
 - Setting Reading Goals Review Questions Thermochem
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Questions Thermochem
 - Fact-Checking eBook Content of Review Questions Thermochem
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- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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