

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Instructions

Standard Conditions

Unit	Standard Pressure
atm	1.013
torr	760
mm	760
cm	760
mm Hg	760

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Gas Law	Equation
Boyle's Law	$P_1V_1 = P_2V_2$
Charles's Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Ideal Gas Law	$PV = nRT$

The constant R (used in the ideal gas law) depends on the unit of pressure:

- if atm, $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$
- if mm, $R = 62.36 \frac{\text{L} \cdot \text{mm Hg}}{\text{mol} \cdot \text{K}}$

Notes:

- Temperature must always be in Kelvin! $K = ^\circ C + 273$
- If you are given two different units of pressure, convert (mm).
- Create two columns before solving: initial and final. Write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you're given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (Mol/L)
- Combined Gas Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

Charles' Law $V_1/T_1 = V_2/T_2$

1. The temperature under my refrigerator is about 40°C. The gas in the refrigerator is at 40°C. The temperature of the refrigerator is 40°C. The temperature of the refrigerator is 40°C. The temperature of the refrigerator is 40°C.

Gay-Lussac's Law $P_1/T_1 = P_2/T_2$

1. A gas is contained in a cylinder at a pressure of 1.0 atm and a temperature of 27°C. The gas is then heated to 127°C. What is the new pressure?

Boyle's Law $P_1V_1 = P_2V_2$

1. A gas is contained in a cylinder at a pressure of 1.0 atm and a volume of 2.0 L. The gas is then compressed to a volume of 1.0 L. What is the new pressure?

Avogadro's Law $V_1/n_1 = V_2/n_2$

1. A gas is contained in a cylinder at a volume of 2.0 L and contains 1.0 mol of gas. The gas is then heated to a volume of 4.0 L. How many moles of gas are now in the cylinder?

Combined Gas Law $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

1. A gas is contained in a cylinder at a pressure of 1.0 atm, a volume of 2.0 L, and a temperature of 27°C. The gas is then heated to a pressure of 2.0 atm and a temperature of 127°C. What is the new volume?

Ideal Gas Law $PV = nRT$

1. A gas is contained in a cylinder at a pressure of 1.0 atm, a volume of 2.0 L, and a temperature of 27°C. How many moles of gas are in the cylinder?

All Gas Laws Included!

Review Packet Gas Laws Answer

Terry C. Jones

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Review Packet Gas Laws Answer:

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Table of Contents Review Packet Gas Laws Answer

1. Understanding the eBook Review Packet Gas Laws Answer
 - The Rise of Digital Reading Review Packet Gas Laws Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Packet Gas Laws Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Packet Gas Laws Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Packet Gas Laws Answer
 - Personalized Recommendations
 - Review Packet Gas Laws Answer User Reviews and Ratings
 - Review Packet Gas Laws Answer and Bestseller Lists
5. Accessing Review Packet Gas Laws Answer Free and Paid eBooks

- Review Packet Gas Laws Answer Public Domain eBooks
- Review Packet Gas Laws Answer eBook Subscription Services
- Review Packet Gas Laws Answer Budget-Friendly Options
- 6. Navigating Review Packet Gas Laws Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Review Packet Gas Laws Answer Compatibility with Devices
 - Review Packet Gas Laws Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Packet Gas Laws Answer
 - Highlighting and Note-Taking Review Packet Gas Laws Answer
 - Interactive Elements Review Packet Gas Laws Answer
- 8. Staying Engaged with Review Packet Gas Laws Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Packet Gas Laws Answer
- 9. Balancing eBooks and Physical Books Review Packet Gas Laws Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Packet Gas Laws Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Packet Gas Laws Answer
 - Setting Reading Goals Review Packet Gas Laws Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Packet Gas Laws Answer
 - Fact-Checking eBook Content of Review Packet Gas Laws Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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