

CHAPTER 24: NUCLEAR

Nuclear Chemistry**Second Half-Year**

Section 1: Write the balanced nuclear equation for the following reactions in the space provided.

1. The radioactive isotope ^{238}Pu decays by emitting an alpha particle. Write the balanced nuclear equation for this reaction. Identify the daughter nucleus. (The atomic number of ^{238}Pu is 94.)
- _____
- _____

2. In a typical nuclear power plant, the reaction $^{235}\text{U} + \text{neutron} \rightarrow ^{236}\text{U} + \text{neutron} + \text{energy}$ is used. Write the balanced nuclear equation for this reaction. Identify the daughter nucleus. (The atomic number of ^{235}U is 92.)
- _____
- _____

3. Write the balanced nuclear equations for the following reactions:

a. $^{238}\text{Pu} \rightarrow \text{alpha particle} + \text{daughter nucleus}$

b. $^{235}\text{U} + \text{neutron} \rightarrow ^{236}\text{U} + \text{neutron} + \text{energy}$

c. $^{238}\text{Pu} \rightarrow \text{beta particle} + \text{daughter nucleus}$

4. Write the complete nuclear equations for the following reactions:

a. $^{238}\text{Pu} \rightarrow \text{alpha particle} + \text{daughter nucleus}$

b. $^{235}\text{U} + \text{neutron} \rightarrow ^{236}\text{U} + \text{neutron} + \text{energy}$

c. $^{238}\text{Pu} \rightarrow \text{beta particle} + \text{daughter nucleus}$

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J Spring



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