

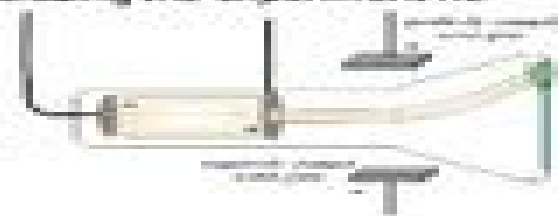
Atomic Theory & Structure: Atomic Theory & Isotope Skill Review

1) State the five parts to Dalton's Atomic Theory.

1. All elements are made of atoms
2. All atoms of an element are identical
3. All atoms of different elements are different
4. Law of constant composition
5. Law of conservation of matter

2) Describe J.J. Thomson's theory of atomic structure including the experiment he performed to reach this theory.

After conducting his research using a gas-filled cathode ray tube where Thomson pumped electricity through the gas sample, he was able to conclude that atoms release electron, contains negatively charged particles called electrons.



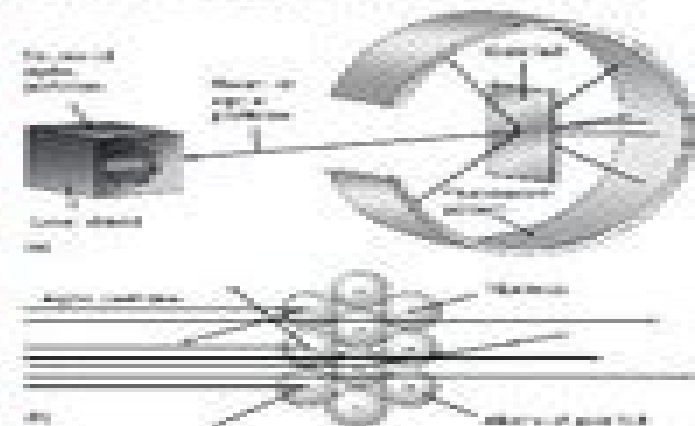
3) Describe J.J. Thomson's theory of atomic structure. Include a picture.

Atoms are a cloud of positive charges with electrons randomly distributed throughout.



4) Describe Ernest Rutherford's theory of atomic structure including the experiment he performed to reach this theory. Include a picture.

Rutherford conducted the gold-foil experiment to confirm the Thomson's theory of atomic structure. The experiment was done with a radioactive source of alpha particles aimed at a piece of gold foil with a detection screen to locate positively charged alpha particles either when they pass through the gold atoms or are deflected. The results showed that most of the alpha particles were detected behind the foil and only a small percentage deflected. After calculating deflection patterns, the only alpha particles that deflected were those that came close to the center of the atom. Based on the results, Rutherford made three conclusions: 1) Plum pudding model is incorrect, 2) the atom has a positively charged center, 3) the rest of the atom (where electrons are located) is empty space.



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