

Chapter 1

Gases and the Zeroth Law of Thermodynamics

- 1.2.** A system is any part of the universe under observation. The “surroundings” includes everything else in the universe. Consider a solution calorimeter in which two aqueous solutions are mixed and temperature changes are recorded. In this case the solutes (reactants and products) would be considered the “system”. The water, calorimeter, and rest of the lab and world would be the “surroundings”.
- 1.4.** (a) $12.56 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 12,560 \text{ cm}^3 = 1.256 \times 10^4 \text{ cm}^3$
(b) $45^\circ\text{C} + 273.15 = 318 \text{ K}$
(c) $1.055 \text{ atm} \times \frac{1.01325 \text{ bar}}{1 \text{ atm}} \times \frac{100,000 \text{ Pa}}{1 \text{ bar}} = 1.069 \times 10^5 \text{ Pa}$
(d) $1233 \text{ mmHg} \times \frac{1 \text{ torr}}{1 \text{ mmHg}} \times \frac{1 \text{ atm}}{760 \text{ torr}} \times \frac{1.01325 \text{ bar}}{1 \text{ atm}} = 1.644 \text{ bar}$
(e) $125 \text{ mL} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 125 \text{ cm}^3$
(f) $4.2 \text{ K} - 273.15 = -269.0^\circ\text{C}$
(g) $25750 \text{ Pa} \times \frac{1 \text{ bar}}{100,000 \text{ Pa}} = 0.2575 \text{ bar}$
- 1.6.** $p_{\text{tot}} = p_{\text{atmosph}} + \rho hg$, where ρhg correspond to the pressure exerted by the liquid
 $p_{\text{tot}} - p_{\text{atmosph}} = \rho hg = (1.0 \times 10^3 \text{ kg/m}^3)(0.23 \text{ m})(9.80 \text{ m/s}^2) = 2254 \text{ N/m}^2 = 2.3 \times 10^3 \text{ Pa}$
- 1.8.** In terms of the zeroth law of thermodynamics, heat will flow from the (hot) burner or flame on the stove into the (cold) water, which gets hotter. Then heat will move from the hot water into the (colder) egg.
- 1.10.** For this sample of gas under these conditions, $F(T) = 2.97 \text{ L} \times 0.0553 \text{ atm} = 0.164 \text{ L} \cdot \text{atm}$. If the pressure were increased to 1.00 atm: $0.164 \text{ L} \cdot \text{atm} = (1.00 \text{ atm}) \times V$; therefore $V = 0.164 \text{ L}$.
- 1.12.** $V = R \cdot \frac{nT}{p}$, which rearranges to $R = \frac{pV}{nT}$.
Therefore: $R = \frac{(2.66 \text{ bar})(27.5 \text{ L})}{(1.887 \text{ mol})(466.9 \text{ K})} = 0.0830 \frac{\text{L} \cdot \text{bar}}{\text{mol} \cdot \text{K}}$

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Table of Contents Physical Chemistry David W Ball Solution Manual

1. Understanding the eBook Physical Chemistry David W Ball Solution Manual
 - The Rise of Digital Reading Physical Chemistry David W Ball Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry David W Ball Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Chemistry David W Ball Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry David W Ball Solution Manual
 - Personalized Recommendations
 - Physical Chemistry David W Ball Solution Manual User Reviews and Ratings
 - Physical Chemistry David W Ball Solution Manual and Bestseller Lists

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 - Highlighting and Note-Taking Physical Chemistry David W Ball Solution Manual
 - Interactive Elements Physical Chemistry David W Ball Solution Manual
8. Staying Engaged with Physical Chemistry David W Ball Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Chemistry David W Ball Solution Manual
9. Balancing eBooks and Physical Books Physical Chemistry David W Ball Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Chemistry David W Ball Solution Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Chemistry David W Ball Solution Manual
 - Setting Reading Goals Physical Chemistry David W Ball Solution Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Chemistry David W Ball Solution Manual
 - Fact-Checking eBook Content of Physical Chemistry David W Ball Solution Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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