

# OCR: Chemistry – Unit 2 – Chains, Energy and Resources

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**Note: For all Mechanisms and Reactions – There is an attached sheet at the end showing all the needed reactions in more detail.**

THERE IS ALSO A LIST OF ALL EQUATIONS AND REACTION NEEDED FOR THIS SYLLABUS ATTACHED IN A SEPARATE FILE.

(a) Interpret and use the terms:

(i) Empirical formula as the **simplest whole number ratio of atoms of each element present in a compound.**

(ii) molecular formula as the **actual number of atoms of each element in a molecule,**

(iii) general formula as the **simplest algebraic formula of a member of a homologous series:**

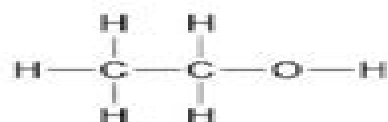
|             |                 |
|-------------|-----------------|
| Alkane      | $C_nH_{2n+2}$   |
| Alkene      | $C_nH_{2n}$     |
| Alcohol     | $C_nH_{2n+2}OH$ |
| Cycloalkane | $C_nH_{2n}$     |
| Cycloalkene | $C_nH_{2n-2}$   |
| Haloalkane  | $C_nH_{2n+2}X$  |

(iv) structural formula as **the minimal detail that shows the arrangement of atoms in a molecule**

E.g., Butane:

$CH_3CH_2CH_2CH_3$  or  $CH_3(CH_2)_2CH_3$ ,

(v) displayed formula as **the relative positioning of atoms and the bonds between them, ie for ethanol:**



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**LP Steffe**



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