

# of substance:

## Substance:

Mass of particles of a substance  
in mol

$6.02 \times 10^{23}$  particles  
amount of substance

## Constant:

particles per mol,  $6.02 \times 10^{23} \text{ mol}^{-1}$

## Unit:

mol, measured in g  
equivalent to Mr

## Atomic Structure

### Isotopes:

→ Same # of protons, same element  
of p<sup>+</sup> and e<sup>-</sup> but different  
neutrons different mass

### Physical Properties:

→ Can be different - e.g. different  
melting + boiling points  
→ i.e. heavy water has higher melting  
point + boiling point + density

### Chemical Properties:

→ Same no. of electrons + same electronic  
structure in outer shell + transfer of e<sup>-</sup> is  
unaffected hence have same chemical  
properties.  
→ Different no. of n has no effect  
on bonding + reactions

### Compounds + Ionic Formulae:

Alkali:  $\text{NO}_3^-$

### Molar Gas Volume:

→ Volume of gas per mole of gas molecules in a stated temperature + pressure  
- Measured in  $\text{dm}^3 \text{ mol}^{-1}$  or  $\text{m}^3 \text{ mol}^{-1}$

At RTP (20°C and 100kPa):  
(293K and 100000 Pa)

is  $24 \text{ dm}^3 \text{ mol}^{-1}$   
at RTP.

Volume /  
 $\text{dm}^3$

## Electronegativity:

→ the measure of attraction of a bonded atom  
in a covalent bond

### Electronegativity increases across period:

- Increase in no. of protons in an atom  
→ Nuclear charge increases

### Electronegativity DE:

- More electron shells  
→ More shielding  
→ Atomic radius increases  
→ Nuclear charge decreases  
→ Nuclear attraction decreases  
→ Electronegativity decreases

# OCR A LEVEL CHEMISTRY MODULE 2 DIAGRAMS



→ Consists of a small dense  
nucleus of protons + neutrons  
surrounded by electrons orbiting  
in electron shells in shells

- No. of p<sup>+</sup> = No. of e<sup>-</sup>

→ Equal + opposite charges cancel  
in atom has neutral charge  
- Neutrons important in nuclear  
stability - Electrostatic repulsion  
between p<sup>+</sup> in small dense nucleus

### Mass Number:

→ No. of p<sup>+</sup> + n in nucleus of atom  
- p<sup>+</sup> and n have relative mass of 1  
and e<sup>-</sup> mass is negligible is 0

### Relative Isotopic Mass:

→ The mass of an atom of an isotope  
relative to  $1/12$ th the mass of an atom of C-12

### Relative Atomic Mass:

→ The weighted mean mass of an atom of  
an element relative to  $1/12$ th the mass of  
an atom of C-12

### Weighted Mean Mass:

→ Average mass of all isotopes, taking into  
account the relative abundance of each isotope

- It dipoles  
Molecule is  
out + inlets

e.g.  $\text{H}_2\text{O}$

- O is more electronegative than H  
- O-H has permanent dipole  
since do not cancel there is an overall dipole

e.g.  $\text{CO}_2$

O is more electronegative than C  
C=O has a dipole  
but the bond is non-polar

# Ocr Chemistry A Level Unit 2 2014 Paper

**LM Reder**



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