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Preparation of Soluble Salts

By Neutralization

25.0cm² acid, as standard solution, is placed in conical flask using pipette. Add few drops of indicator & titrate with alkali from burette until indicator changes colour, showing all acid has just reacted. Volume of alkali added is measured. Prepare new 25.0cm² acid again with pipette & add same volume of alkali as before to prevent excess alkali/acid because both reactant & product are aqueous. Next, the product is engograped to drive us to obtain the salt.



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By Reacting Metal with Asid





Only metals like sinc and magnesium, which moderately react with dilute acids, are used.

E.a. Reacting In with H.SO, to prepare InSO,

$$2n(ij) + H_iSO_i(aq) \rightarrow 2nSO_i(aq) + H_iO(i)$$

Zn is added to H₂SO₄ until in excess to ensure no more H₂SO₄ is present. Then the maxture is filtered off to separate Zn from ZrSO₄. The filtrate (ZxSO₄) is then placed in evaporating dish to evaporate most of water then it's cooled after ZrSO₄ crystals are formed. The crystals then filtered and squeezed between filter papers to dry.

By Reacting Insoluble Base with Acid

E.g. Reacting MgO with Acids

The same method as reaction of acid with metal is used, so refer to diagram and above explanation, substituting reactants and products.

By Reacting Carbon ate with Acid

E.g. Reacting CoCO, with Acids

$$K_1CO_1(x) + H_1SO_2(xx) \rightarrow K_2SO_2(xx) + CO_2(x) + H_1O(x)$$

The same process is used as reaction of acid with metal, just that carbon dioxide is produced. Carbon dioxide can be tested by bubbling it into limewater which will turn limewater colouriess to milky.

7.4 Properties and Uses of Ammonia

Ammonia and its Uses

Are produced from nitrogen reacted with hydrogen.

for producing: fertilisers, nitric sold, nylon, dyes, cleaners and dry cell

The Manufacture of Ammonia: The Hober Process

The Process.

Nitrogen and hydrogen are mixed together in ratio 1:3, where nitrogen is obtained from air and hydrogen is obtained from natural gas, and passed over iron-catalvit.



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