



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

Duration: 120
Time: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use permanent marker for writing and drawing.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

- 3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹ and 1 mol of 1-hexanol.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B or of A, B and C or of others as well? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and is the prime reason of their volatility. (2)
- 3.1.4 What factor of the compounds A and B is responsible for the difference in their boiling points? (2)
- 3.1.5 Draw a Lewis dot-and-cross diagram for compound A. (2)
- 3.1.6 Which compound (A or B) will have the higher vapour pressure? (2)
- 3.1.7 Draw a Lewis dot-and-cross diagram for compound B. (2)

Compound C (M_r 102) is shown in the table. Draw a suitable molecular model that will compound A and B for the intermolecular forces existing in it. (2)

- 3.1.8 Explain why the boiling point of compound B is 100°C higher than that of compound A and B by referring to the ΔH_{vap} and intermolecular forces involved. (2)

- 3.2 Substances D and E are shown in the table below. At room temperature (25°C) and 1 atm, D is a solid and E is a liquid. Draw the structural formula of D and E. (2)



- 3.2.2 Explain the reason for the difference in the boiling points of D and E by referring to the intermolecular forces involved. (2)

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