

Topic	Grade	What do you need to know	 or 	What do You have to work on?
Atomic Structure	A	- Why an atom is neutral in terms of sub-atomic particles - Outline isotopes from data about numbers of protons and neutrons		
	B	- Know the charges and masses of protons, electrons and neutrons. - Know how isotopes are the same but have different mass number.		
	C	- Describe the arrangement of an atoms electron structure - 2.8.8.2 - Describe the arrangement of elements in the periodic table.		
Ionic Bonding	A	- Explain ionic bonding using the 'dot and cross' model. - Explain the melting point of NaCl by the structure of and bonding. - Explain why the melting point of NaCl is lower than that of MgO.		
	B	- Predict the properties of substances with giant ionic structures. - Understand that atoms with a full outer shell are electronically stable.		
	C	- Ionic bonds are formed from metals and non-metals - Electrons are donated and received. - Metals always donate the electron(s). - Non- metals always receive the electron(s)		
The Periodic Table	A	- Explain some covalent bonds using a 'dot and cross' model - Explain double covalent binds using a 'dot and cross' model		
	B	Know why CO₂ and H₂O have low melting points		
	C	- Covalent bonds always have non-metals and share electrons - Know that the group number is the same as the number of electrons in the outer shell. - Be able to find out how many shells an atom has from the period it is in.		
The Group 1 Elements	A	- Construct a balanced symbol equation for Cu²⁺, Fe²⁺ and Fe³⁺ to make Cu(OH)₂, Fe(OH)₂, Fe(OH)₃. - Reactivity with water gives NaOH and KOH. - Know the process is oxidation.		
	B	- Why group 1 elements form similar positive ions. - Construct a balance symbol equation of group 1 with water		
	C	- Metal + Water → Metal Hydroxide + Hydrogen - Know how to use a flame test to identify potassium, lithium and sodium compounds.		
The Group 7 Elements	A	- Predict if displacement reactions will occur with other halogens - Explain the trend of reactivity of group 7 by ideas of electron gain - Explain why a process is reduced from its ionic		

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