

Reference Tables for Physical Setting/EARTH SCIENCE

Radioactive Decay Data

RADIOACTIVE ISOTOPE	DISINTEGRATION	HALF-LIFE (years)
Carbon-14	$^{14}\text{C} \rightarrow ^{14}\text{N}$	5.7×10^3
Potassium-40	$^{40}\text{K} \rightarrow ^{40}\text{Ar}$ $^{40}\text{K} \rightarrow ^{40}\text{Ca}$	1.3×10^9
Uranium-238	$^{238}\text{U} \rightarrow ^{206}\text{Pb}$	4.5×10^9
Rubidium-87	$^{87}\text{Rb} \rightarrow ^{87}\text{Sr}$	4.9×10^{10}

Specific Heats of Common Materials

MATERIAL	SPECIFIC HEAT (Joules/gram • °C)
Liquid water	4.18
Solid water (ice)	2.11
Water vapor	2.00
Dry air	1.01
Basalt	0.84
Granite	0.79
Iron	0.45
Copper	0.38
Lead	0.13

Equations

$$\text{Eccentricity} = \frac{\text{distance between foci}}{\text{length of major axis}}$$

$$\text{Gradient} = \frac{\text{change in field value}}{\text{distance}}$$

$$\text{Rate of change} = \frac{\text{change in value}}{\text{time}}$$

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

Properties of Water

Heat energy gained during melting	334 J/g
Heat energy released during freezing	334 J/g
Heat energy gained during vaporization	2260 J/g
Heat energy released during condensation	2260 J/g
Density at 3.98°C	1.0 g/mL

Average Chemical Composition of Earth's Crust, Hydrosphere, and Troposphere

ELEMENT (symbol)	CRUST		HYDROSPHERE	TROPOSPHERE
	Percent by mass	Percent by volume	Percent by volume	Percent by volume
Oxygen (O)	46.10	94.04	33.0	21.0
Silicon (Si)	28.20	0.88		
Aluminum (Al)	8.23	0.48		
Iron (Fe)	5.63	0.49		
Calcium (Ca)	4.15	1.18		
Sodium (Na)	2.36	1.11		
Magnesium (Mg)	2.33	0.33		
Potassium (K)	2.09	1.42		
Nitrogen (N)				78.0
Hydrogen (H)			66.0	
Other	0.91	0.07	1.0	1.0

2011 EDITION

This edition of the Earth Science Reference Tables should be used in the classroom beginning in the 2011–12 school year. The first examination for which these tables will be used is the January 2012 Regents Examination in Physical Setting/Earth Science.

Eurypterus remipes

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Nys Reference Sheet Review Earth

David Kirk



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Table of Contents Nys Reference Sheet Review Earth

1. Understanding the eBook Nys Reference Sheet Review Earth
 - The Rise of Digital Reading Nys Reference Sheet Review Earth
 - Advantages of eBooks Over Traditional Books
2. Identifying Nys Reference Sheet Review Earth
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nys Reference Sheet Review Earth
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nys Reference Sheet Review Earth
 - Personalized Recommendations
 - Nys Reference Sheet Review Earth User Reviews and Ratings
 - Nys Reference Sheet Review Earth and Bestseller Lists
5. Accessing Nys Reference Sheet Review Earth Free and Paid eBooks
 - Nys Reference Sheet Review Earth Public Domain eBooks
 - Nys Reference Sheet Review Earth eBook Subscription Services
 - Nys Reference Sheet Review Earth Budget-Friendly Options
6. Navigating Nys Reference Sheet Review Earth eBook Formats

- ePub, PDF, MOBI, and More
- Nys Reference Sheet Review Earth Compatibility with Devices
- Nys Reference Sheet Review Earth Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nys Reference Sheet Review Earth
 - Highlighting and Note-Taking Nys Reference Sheet Review Earth
 - Interactive Elements Nys Reference Sheet Review Earth
- 8. Staying Engaged with Nys Reference Sheet Review Earth
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nys Reference Sheet Review Earth
- 9. Balancing eBooks and Physical Books Nys Reference Sheet Review Earth
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nys Reference Sheet Review Earth
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nys Reference Sheet Review Earth
 - Setting Reading Goals Nys Reference Sheet Review Earth
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nys Reference Sheet Review Earth
 - Fact-Checking eBook Content of Nys Reference Sheet Review Earth
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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