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Post-Laboratory Report

Experiment No.: 7.1 (Practical Exam)

Experiment Title: The Boiling Reaction

I. INTRODUCTION

A neutralization reaction can be defined as a reaction in which an acid and a base react together to form water and salt as products. In a neutralization reaction, there is a combination of H^+ ions and OH^- ions which form water. This type of reaction is usually an acid-base reaction. The pH of the products varies depending on the strength of the reacting acid and base. A common example of a neutralization reaction includes the formation of sodium chloride (salt).

Neutralization is also very important in our daily lives. For instance, toothpaste contains bases that neutralize the acid that our mouth creates from bacteria and to make the cakes grow, baking powder is used. In addition, farmers also use bases in order to neutralize acidic soils. This is done in order to prevent deficiencies in major plant nutrients since nutrients are less available to plants in acidic soil. An example of a material mixed in the soil to neutralize its acidity is calcium carbonate (limestone). Even in our bodies, neutralization is present. When food is moved between our stomach and intestines, the food needs to be neutralized.

As an application to all the learnings that we got about neutralization, in this experiment, we will be observing how hydrochloric acid and sodium hydroxide will react when they are mixed together. We will also describe the physical changes during a neutralization reaction such as the release of heat.

II. OBJECTIVES

- To define a neutralization reaction
- To identify the names of the substances involved: acid, base, and salt
- To write chemical equations for neutralization reactions
- To recognize that an acid-alkali reaction can be simplified as the reaction of H^+ ions with OH^- ions to form water

Neutralization Study Guide Answers

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