

# Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

**K, Na, O, S, Cl**

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

**Na, K, S, O, Cl**

3. Which atom has a higher first ionization energy than sodium?

**Chlorine (because it has a higher electronegativity and is closer to the nucleus)**

4. Which two elements in the same family generally have similar properties?

**Sulfur and chlorine (because they are both in the same period and have similar properties)**

5. Arrange the following elements in order of increasing atomic radius from left to right: sodium, chlorine, sulfur.

- a. sodium (because it has the lowest electronegativity) **Na**  
 b. sulfur (because it has the highest electronegativity) **S**  
 c. chlorine (because it has the highest electronegativity) **Cl**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest atomic radius?

**Period 7, Francium (because it has the highest electronegativity and is closest to the nucleus)**

7. Which period in the periodic table shows atoms with the greatest electronegativity? Which element has the greatest electronegativity?

**Period 2, Fluorine (because it has the highest electronegativity and is closest to the nucleus)**

8. Write the atomic radius of each of the following elements.

Na	186
Cl	99
S	104
O	73
K	227
Ca	197
Fe	125

Al	143
Si	111
P	106
S	104
Cl	99
Ar	71

9. Write the atomic radius of each of the following elements.

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**R Bogdan**

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## **Table of Contents Physical Science If8767 Pg 63 Key**

1. Understanding the eBook Physical Science If8767 Pg 63 Key
  - The Rise of Digital Reading Physical Science If8767 Pg 63 Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Pg 63 Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science If8767 Pg 63 Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Pg 63 Key
  - Personalized Recommendations
  - Physical Science If8767 Pg 63 Key User Reviews and Ratings
  - Physical Science If8767 Pg 63 Key and Bestseller Lists
5. Accessing Physical Science If8767 Pg 63 Key Free and Paid eBooks

- Physical Science If8767 Pg 63 Key Public Domain eBooks
  - Physical Science If8767 Pg 63 Key eBook Subscription Services
  - Physical Science If8767 Pg 63 Key Budget-Friendly Options
6. Navigating Physical Science If8767 Pg 63 Key eBook Formats
- ePub, PDF, MOBI, and More
  - Physical Science If8767 Pg 63 Key Compatibility with Devices
  - Physical Science If8767 Pg 63 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science If8767 Pg 63 Key
  - Highlighting and Note-Taking Physical Science If8767 Pg 63 Key
  - Interactive Elements Physical Science If8767 Pg 63 Key
8. Staying Engaged with Physical Science If8767 Pg 63 Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science If8767 Pg 63 Key
9. Balancing eBooks and Physical Books Physical Science If8767 Pg 63 Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science If8767 Pg 63 Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physical Science If8767 Pg 63 Key
- Setting Reading Goals Physical Science If8767 Pg 63 Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science If8767 Pg 63 Key
- Fact-Checking eBook Content of Physical Science If8767 Pg 63 Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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