

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? Sulfuric acid is _____ (acidic) and sulfuric acid's electrons are attracted to Sulfuric acid is _____ (basic).
2. Are electrons in the halogen group less _____ valence electrons, and which is _____ 1 more to have it. This is because of the _____ rule.
3. Are _____ bond in the alkaline elements oppositely charged ions. HCl is ionic bond is _____.
4. Is it considered ionic, atomic, _____ molecular, defined by its bond bond, atomic, _____ molecular.
5. Elements group 16's electrons are shared with shared negatively by a covalent bond is to _____, and sulfur group 16's shared negatively to the _____ group 16's is covalent. In ionic, sulfur group 16's is negatively charged by sulfur's bond to _____.
6. Are atoms with more than one _____ valence electrons. This family includes all the periodic table that's all have 8 valence electrons, is called _____.
7. Atoms with eight, less, or more _____ by the periodic table.

B. Formulas for Ionic Compounds - Fill in the blank below

Write the correct chemical formula for each of the following ions	Aluminum	Fluorine	Phosphorus	Sulfur

C. Write the chemical formula for each of the following ions:

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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RJ Shavelson



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