

# CRUDE OIL

## Task 1 - Crude oil key word definitions:

Crude oil is a finite resource found in rocks. It is the remains of plant & animal life that was buried in mud many millions of years ago.

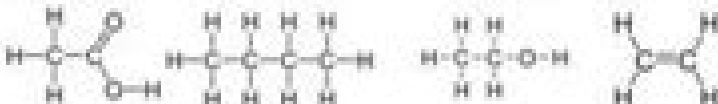
a) Define finite resource

Crude oil is a mixture of a very large number of compounds. Most of the compounds in crude oil are hydrocarbons.

Define Hydrocarbon

## Task 2 - Hydrocarbons

a) Highlight which of the following are not hydrocarbons. Explain your answer



## Task 3 - Alkanes

Alkanes are the simplest type of hydrocarbon. They contain only single bonds & have the general formula:  $\text{C}_n\text{H}_{2n+2}$ . Complete the following table of alkanes:

| Name    | Molecular formula      | Displayed formula                                                                           |
|---------|------------------------|---------------------------------------------------------------------------------------------|
| methane | $\text{CH}_4$          | $\begin{array}{c} \text{H} \\   \\ \text{H}-\text{C}-\text{H} \\   \\ \text{H} \end{array}$ |
| _____   | $\text{C}_2\text{H}_6$ |                                                                                             |
| propane | $\text{C}_3\text{H}_8$ |                                                                                             |
| butane  |                        |                                                                                             |

## Task 4 - Fractional distillation:

a) What is the process of fractional distillation used for?

b) Add labels to the diagram opposite to make clear what happens as you go up the column to the following:

- Size of molecule
- Boiling point of hydrocarbon
- Viscosity of hydrocarbon
- Flammability

c) What property of the molecules does fractional distillation rely upon to separate them from crude oil?

d) Bitumen is collected at the bottom of the column. It is still a liquid & did not turn into a vapour when heated. What does that tell you about the boiling point of bitumen compared to the rest of the hydrocarbons in crude oil?

e) Choose the correct word from the word bank to complete this description of the process of fractional distillation:

Fractional distillation separates the crude oil into different \_\_\_\_\_ according to the \_\_\_\_\_ of the different hydrocarbons.

So each fraction contains a mixture of different-sized hydrocarbons with similar boiling points.

The longer molecules will have greater \_\_\_\_\_ of attraction between them & so will need more \_\_\_\_\_ to overcome these forces. This means these \_\_\_\_\_ chain hydrocarbons will boil at \_\_\_\_\_ temperatures.

|                       |        |           |
|-----------------------|--------|-----------|
| Intermolecular forces | higher | fractions |
| boiling points        | energy | longer    |

## Task 5 - Combustion:

Hydrocarbons are used as fuels as they release a lot of energy when they are burnt.

a) Write the word equation to show the combustion of a hydrocarbon fuel:

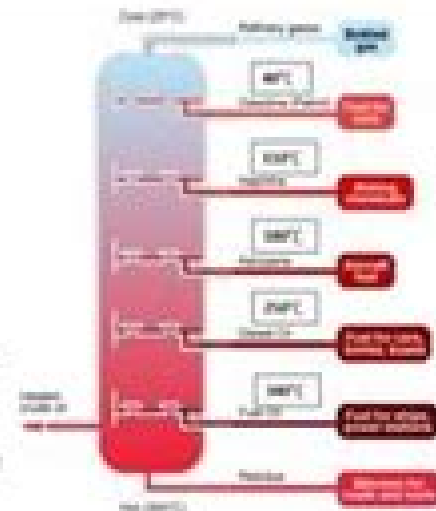
HYDROCARBON + \_\_\_\_\_ → \_\_\_\_\_ + \_\_\_\_\_

b) Following is the equation for the combustion of methane gas, CAN YOU BALANCE IT?



c) What do we call the reaction that happens when we try to burn a fuel without enough oxygen?

d) What is the toxic gas that is produced by this chemical reaction? Explain why it is harmful.



# Ocr Gcse Chemistry C7 Mark Scheme

**E Durkheim**



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