

A reaction in which the Products can react to form the reactants is called a reversible reaction.



Chemical equilibrium occurs when the reactants in a reversible reaction form products at the same rate that products form reactants.

At equilibrium:

- the concentration of the reactants and products does not change.
- the concentration of reactants can be equal to, greater, or less the concentration of the products.



$$K_{eq} = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

If K_{eq} is	reaction is favored
= 1	neither
< 1	reverse
> 1	forward

Ex. #1: Calculate K_{eq} for the following reaction using the given equilibrium concentrations. Then determine whether the forward or **reverse** reaction is favored.



$$K_{eq} = 6.0 \times 10^{-3}$$

Note Taking Guide Episode 1202 Answers

Rachel Sandford



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