

# THERMOCHEMICAL EQUATIONS AND TYPES OF ENTHALPY CHANGES



## Thermochemical Equations

A thermochemical equation is a reaction equation accompanied by an enthalpy change ( $\Delta H$ ). In addition to stating the number of moles reactants and the number of moles product, it also displays the amount of heat released or absorbed in the reaction in units of kJ.

### Example :

The thermochemical equation is known as:



This means that the formation of 1 mol  $\text{H}_2\text{O}$  hydrogen gas and oxygen gas released energy of 285,85 kJ (exothermic reaction).

Some things that must be considered in the thermochemical equation are

### 1. THE REACTION COEFFICIENT EXPRESSES MOLES.

#### Example:



(means that when 1 mol  $\text{H}_2$  reacts with 1 mol  $\text{I}_2$  to produce 2 mol  $\text{HI}$ , heat is absorbed by 51.9 kJ).

### 2. IF THE REACTION IS REVERSED, $\Delta H$ CHANGES SIGN

From reaction 1:



### 3. IF THE REACTIONS ARE MULTIPLIED, THEN $\Delta H$ IS MULTIPLIED

From reaction 1:



# Notetaking Guide For Thermochemical Equations

**Michael Brown**



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Thermochemistry is the science of analyzing molecular reactions to determine if they are spontaneous energy absorbing or releasing and to predict the product mole ratios and rates Chemical reactions like most other processes tend to follow the path of free energy minimization or entropy maximization This principle forms the mathematical basis for the analytical approach This book is a how to manual filled with many examples and comes with all the code you need to accomplish this task      *The Ultimate Chemical Equations Handbook, Teacher Edition* Jane D. Smith,George R. Hague,2011

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