

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- Boyle's Law

Boyle's Theory and Boyle's Law

- Boyle's of gases is caused by the elasticity of the gas molecules themselves that collide.
- If the gas is compressible to small the volume is small which increases the pressure which is small.
- Force exerts pressure on closed on the walls of the container.
- Strength of the gas will increase.

- Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. After the volume is reduced, it maintains a volume of 400 ml. above the pressure is 1000 mm. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad 1000 \text{ mm} \times 400 \text{ ml} = P_2 \times 800 \text{ ml} = P_2$$

$$P_2 = \frac{1000 \times 400}{800} = 500 \text{ mm}$$

- Ex. 2. A gas is pressure of 100 mm. and 400 ml. of gas has a volume of 200 ml. If the pressure is increased until the volume is 100 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad 100 \text{ mm} \times 400 \text{ ml} = P_2 \times 100 \text{ ml} = P_2$$

$$P_2 = 400 \text{ mm}$$

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