

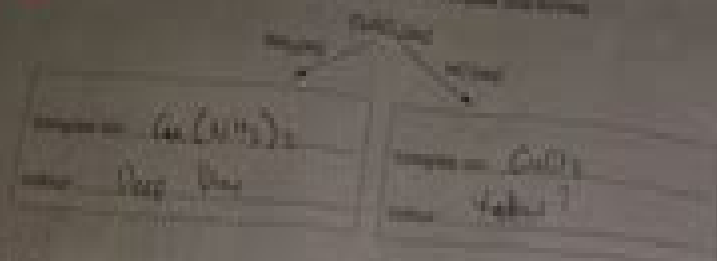
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a copper complex.

In the boxes, write the formula and name of the ligands and metal.



(b) Cu^{2+} has two d-orbitals in d_{sp^2} hybridization state and two water molecules. The d_{sp^2} hybridization is a tetrahedral ligand.

The resulting formula of the complex is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term tetrahedral ligand?

A ligand that forms four bonds with central metal ion.

(d) The reaction is shown in Fig. 1.2 below.

The reaction is about the synthesis of a copper complex. The charge has been given.



Complex A

Complete the 3D diagram of the other two geometries of A. You do not need to include any charges.

Indicate with ticks whether the stereoisomers are cis, trans, optical or a combination of these types.

Stereoisomer														
Type	<table><tr><td>cis</td><td></td></tr><tr><td>trans</td><td>✓</td></tr><tr><td>optical</td><td></td></tr></table>	cis		trans	✓	optical		<table><tr><td>cis</td><td>✓</td></tr><tr><td>trans</td><td></td></tr><tr><td>optical</td><td>✓</td></tr></table>	cis	✓	trans		optical	✓
cis														
trans	✓													
optical														
cis	✓													
trans														
optical	✓													

(e) What is the empirical formula, including the charge, of the complex in A?

$[Cu(NH_3)_2(ox)]^{2+}$

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ML Morrison



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