

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the state of dynamic equilibrium involving interconverting reactants simultaneously forming as reactants or products and reactants. These reactions are called reversible reactions, as the reaction can proceed in either direction. Once the equilibrium is reached the rates of the forward and reverse reactions are equal, and the concentrations of the reactants and products remain constant. The importance of understanding that equilibrium is a dynamic state is that it allows us to predict the behavior of a system. For instance, if a reaction is at equilibrium, then the forward and reverse reactions continue to occur over time, but there is no net change in the concentrations of either the reactants or the products. Equilibrium conditions are reached with an equilibrium constant, or equilibrium constant, which is a ratio of the concentrations of the products to the reactants. Le Chatelier's Principle states that applying stress changes the equilibrium of a system. Le Chatelier's Principle states that if a system is at equilibrium, then any change in the conditions of the system will cause the system to shift in a direction that relieves the change and allow the system to return to equilibrium. When more products are added, an equilibrium reaction will shift to the left to consume the excess products and return to equilibrium. When more reactants are added, the equilibrium will shift to the right to consume the excess reactants and return to equilibrium. In this lab, you will be observing the changes in equilibrium when the amount of reactants, and temperature are changed. Follow Le Chatelier's Principle in the following reactions. Shifted equilibrium is indicated by the color change in solution according to the color indicators in the equations below. For example, if the color changes to red, then reaction shifts to the right to consume the excess reactants and return to equilibrium.



Adding NaOH solution to $\text{Fe}^{3+}(\text{aq})$ and $\text{SCN}^{-}(\text{aq})$ will allow Fe^{3+} ions to react with OH^{-} ions, leaving the percent concentration of Fe^{3+} ions to produce more FeSCN^{2+} ions. Le Chatelier's Principle states that if a system is at equilibrium, then any change in the conditions of the system will cause the system to shift in a direction that relieves the change and allow the system to return to equilibrium.

Procedure (This Group):

- | | |
|---|--|
| 1. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 2. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 3. Prepare 100 mL of 0.1 M NaOH solution. | 4. Prepare 100 mL of 0.1 M HCl solution. |
| 5. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 6. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 7. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 8. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 9. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 10. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 11. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 12. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 13. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 14. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
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| 17. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 18. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 19. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 20. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 21. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 22. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 23. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 24. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 25. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 26. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 27. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 28. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
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| 97. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 98. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 99. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 100. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |

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