

of substance:

Substance:

Mass of particles of a substance
in mol

6.02×10^{23} particles
amount of substance

Constant:

particles per mol, $6.02 \times 10^{23} \text{ mol}^{-1}$

Unit:

mol, measured in g
equivalent to mol

Atomic Structure

Isotopes:

→ Same # of protons, same element
of p^+ and e^- but different
neutrons different mass

Physical Properties:

→ Can be different - e.g. different
melting + boiling points
→ i.e. heavy water has higher melting
point + boiling point + density

Chemical Properties:

→ Same no. of electrons + same electronic
structure in outer shell + transfer of e^- is
unaffected hence have same chemical
properties.
→ Different no. of n has no effect
on bonding + reactions

Compounds + Ionic Formulas:

Alkali NO_3^-

Molar Gas Volume:

→ Volume of gas per mole of gas molecules in a stated temperature + pressure
- Measured in $\text{dm}^3 \text{ mol}^{-1}$ or $\text{m}^3 \text{ mol}^{-1}$

At RTP (20°C and 100kPa):
(293K and 100000Pa)

is $24\text{ dm}^3 \text{ mol}^{-1}$
at RTP.

Volume V
[dm^3]

Electronegativity:

→ the measure of attraction of a bonded atom
in a covalent bond

Electronegativity increases across period:

- Increase in no. of protons in an atom
→ Nuclear charge increases

Electronegativity DE:

- More electron shells
→ More shielding
→ Atomic radius increases
→ Nuclear charge decreases
→ Nuclear attraction decreases
→ Electronegativity decreases

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→ Consists of a small dense
nucleus of protons + neutrons
surrounded by electrons orbiting
in electron shells in shells

- No. of p^+ = No. of e^-

→ Equal + opposite charges cancel
in atom has neutral charge
- Neutrons important in nuclear
stability - Electrostatic repulsion
between p^+ in small dense nucleus

Mass Number:

→ No. of p^+ + n in nucleus of atom
- p^+ and n have relative mass of 1
and e^- mass is negligible is 0

Relative Isotopic Mass:

→ The mass of an atom of an isotope
relative to $1/12$ th the mass of an atom of ^{12}C

Relative Atomic Mass:

→ The weighted mean mass of an atom of
an element relative to $1/12$ th the mass of
an atom of ^{12}C

Weighted Mean Mass:

→ Average mass of all isotopes, taking into
account the relative abundance of each isotope

- It dipoles
Molecule is
out + inlets

e.g. H_2O

- O is more electronegative than H
- O-H has permanent dipole
since do not cancel there is an overall dipole

e.g. CO_2

O is more electronegative than C
 $\text{C}=\text{O}$ has a dipole
but the bond is non-polar

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