

# STAAR CHEMISTRY REFERENCE MATERIALS



## ATOMIC STRUCTURE

$$\text{Speed of light} = (\text{frequency})(\text{wavelength}) \qquad c = f\lambda$$

$$\text{Energy} = (\text{Planck's constant})(\text{frequency}) \qquad E_{\text{photon}} = hf$$

$$\text{Energy} = \frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})} \qquad E_{\text{photon}} = \frac{hc}{\lambda}$$

## BEHAVIOR OF GASES

$$\text{Total pressure of a gas} = \left( \begin{array}{c} \text{sum of the partial pressures} \\ \text{of the component gases} \end{array} \right) \qquad P_T = P_1 + P_2 + P_3 + \dots$$

$$(\text{Pressure})(\text{volume}) = (\text{moles})(\text{ideal gas constant})(\text{temperature}) \qquad PV = nRT$$

$$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})} \qquad \frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

$$(\text{Initial pressure})(\text{initial volume}) = (\text{final pressure})(\text{final volume}) \qquad P_1V_1 = P_2V_2$$

$$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})} \qquad \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})} \qquad \frac{V_1}{n_1} = \frac{V_2}{n_2}$$

## SOLUTIONS

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}} \qquad M = \frac{\text{mol}}{\text{L}}$$

$$\text{Ionization constant of water} = \left( \begin{array}{c} \text{hydrogen ion} \\ \text{concentration} \end{array} \right) \left( \begin{array}{c} \text{hydroxide ion} \\ \text{concentration} \end{array} \right) \qquad K_w = [\text{H}^+][\text{OH}^-]$$

$$\left( \begin{array}{c} \text{Volume of} \\ \text{solution 1} \end{array} \right) \left( \begin{array}{c} \text{molarity of} \\ \text{solution 1} \end{array} \right) = \left( \begin{array}{c} \text{volume of} \\ \text{solution 2} \end{array} \right) \left( \begin{array}{c} \text{molarity of} \\ \text{solution 2} \end{array} \right) \qquad V_1M_1 = V_2M_2$$

$$\text{pH} = -\log(\text{hydrogen ion concentration}) \qquad \text{pH} = -\log[\text{H}^+]$$

## THERMOCHEMISTRY

$$\text{Heat gained or lost} = (\text{mass}) \left( \begin{array}{c} \text{specific} \\ \text{heat} \end{array} \right) \left( \begin{array}{c} \text{change in} \\ \text{temperature} \end{array} \right) \qquad Q = mc_p\Delta T$$

$$\text{Enthalpy of reaction} = \left( \begin{array}{c} \text{enthalpy} \\ \text{of products} \end{array} \right) - \left( \begin{array}{c} \text{enthalpy} \\ \text{of reactants} \end{array} \right) \qquad \Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

# Sample For Staar Testing Questions For Chemistry

**Michael Seilmaier**



## **Sample For Staar Testing Questions For Chemistry:**

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In a world taken by screens and the ceaseless chatter of instantaneous interaction, the melodic splendor and mental symphony created by the published term often fade into the back ground, eclipsed by the constant noise and interruptions that permeate our lives. But, located within the pages of **Sample For Staar Testing Questions For Chemistry** a charming literary value overflowing with raw thoughts, lies an immersive symphony waiting to be embraced. Constructed by an elegant musician of language, that fascinating masterpiece conducts visitors on a mental trip, well unraveling the hidden melodies and profound influence resonating within each carefully constructed phrase. Within the depths of the poignant evaluation, we shall examine the book is key harmonies, analyze their enthralling writing fashion, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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