Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium Chapters 17 & 18

Essential Skills/ State Standards:

9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)

9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)

9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat (q= m · c · Δ T) & latent heat of phase change (q= m Δ H). (CSS: 7d)

9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)

9.5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)

9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS:

9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vecabulary.

- Condensation
- Freezing
- Vaporization

Sublimation.

- Miching
- Bailing
- Condensation

- Deposition
- Heat
- Temperature
- Thermochemistry
- Exathermic
- Endothermic

- Colorimeter
- Joule
- Specific heat
- q=m-c-∆T
- q= m A H assertantes.
- q=m ΔH total

- Heat of reaction
- Le Chatlier's principle
- Equilibrium
- Activation Energy
- Cohelyst

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

- According to the collision theory, what 3 things must happen in order for two molecules to react?
 Must collide with each other, Must collide with enough energy. & in the correct orientation
- What is the importance of the activation energy in a chemical reaction? Amount of energy needed for a reaction to occur.
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:
 HCl(aq) + Mg(s) → MgCl₂(aq) + H₂(g)
 - a. Increase temp: Increase, molecules collide more
 - b. Increase HCI: Increase, more collisions because more reactants to collide.

c. Increase pressure: Hard to say, wouldn't affect the reactants directly because none of them are gases.
One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call., (this one won't be on the test).

d. Remove MgCl2: Increase, less products means less chance for them to interfere with reactant collisions

e. Add H₂: Decrease, more products mean they interfere more with reactants and the reactants would collide less.

- Identify the following terms with their correct letter from the graphs:
 - * Reactants- 1

- * Products- 2
- Activation energy (number/ letter & amount)-

A (300 Joules)

" Is this reaction exothermic or endothermic?

Endothermic

 Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph) Speeds it up by lowering the activation energy.



Reaction Rates And Equilibrium Test A Answers

Katrin Zwirglmaier

Reaction Rates And Equilibrium Test A Answers:

Chemistry Rob Ritchie, 2008-07 Divided into seven manageable day sections this timed revision programme covers essential GCSE topics in double page spreads These spreads indicate how much time should be spent on each section and combine clear and concise explanations flow charts spidergrams and illustrations with progress check questions and answers

Access to Chemistry Avril Higton, Mike Clemmet, Elaine Golding, Alan V Jones, 2007-10-31 This text has been specifically designed to prepare people with previously limited chemical knowledge for entrance into science related courses such as Foundation and Access courses which involve chemistry in higher education Until now there have been no texts available for use on these courses and this book fills that gap Access to Chemistry effectively forms a self study course which is split into separate modules and units covering the full spread of concepts required for those needing a basic knowledge of chemistry The material is presented in a friendly and easy to use manner which allows the student to pace their acquisition of knowledge and gain increasing confidence in order to succeed in understanding essential relevant concepts Other useful features of this book include starter diagnostic tests worked examples and self study tests with answers at the end of each unit In addition to Access or Foundation course students and their tutors to whom this book will prove essential it will have an appeal also as a revision text for those needing a refresher after a break in the subject In addition it will be of interest to members of the general public who wish to better educate themselves on chemical matters as it provides a clear and useful insight into areas such as health home chemicals business market trends and gardening *Cracking the SAT Chemistry* Subject Test, 2013-2014 Edition Princeton Review, 2013-04-23 If you need to know it it s in this book This eBook version of the 2013 2014 edition of Cracking the SAT Chemistry Subject Test has been optimized for on screen viewing with cross linked guestions answers and explanations It includes 3 full length practice tests with detailed explanations Review of all essential content from chemical equations to kinetics to electron configurations Helpful study lists of key lab equipment and a cheat sheet of important equations Key strategies that will help maximize your score Tons of sample problems and drills with detailed explanations Cracking the SAT Chemistry Subject Test, 15th Edition Princeton Review, 2015-02-17 EVERYTHING YOU NEED TO HELP SCORE A PERFECT 800 Equip yourself to ace the SAT Chemistry Subject Test with The Princeton Review's comprehensive study guide including 3 full length practice tests thorough reviews of key chemistry topics and targeted strategies for every question type This eBook edition has been optimized for on screen viewing with cross linked questions answers and explanations We don't have to tell you how tough SAT Chemistry is or how helpful a stellar exam score can be for your chances of getting into your top choice college Written by the experts at The Princeton Review Cracking the SAT Chemistry Subject Test arms you to take on the test and achieve your highest score Techniques That Actually Work Tried and true strategies to help you avoid traps and beat the test Tips for pacing yourself and guessing logically Essential tactics to help you work smarter not harder Everything You Need to Know to Help Achieve a High Score

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CliffsNotes AP Chemistry Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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