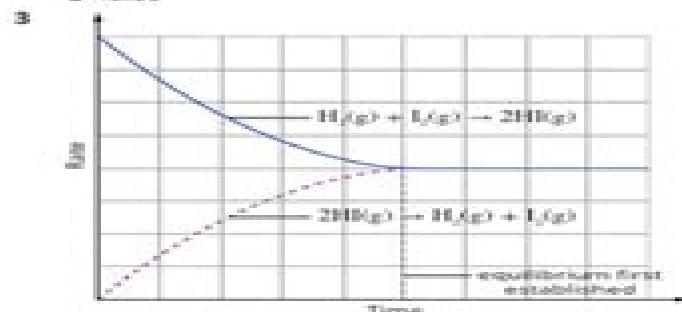


## Chapter 2 Chemical equilibrium

### 2.1 Dynamic equilibrium

#### 2.1 KEY QUESTIONS

1.
  - a An **open** system occurs when both matter and energy are exchanged between the system and the surroundings.
  - b Chemical equilibria can occur in **closed** systems.
  - c In some circumstances, a reversible reaction can form a **dynamic** equilibrium.
  - d The extent of reaction for chemical equilibria is **different** for different equilibria.
2.
  - a false
  - b false
  - c true
  - d false



4. closed, decreases, less, reverse, increases, equal
5.
  - a open; both energy and matter (carbon dioxide and water vapour) are exchanged with the surroundings
  - b closed; only energy can be exchanged with the surroundings. While the can is unopened, no matter is exchanged.
  - c open; both energy and matter are exchanged with the surroundings
6. B and F. Equilibrium can only occur in a closed system. At the molecular level, some of the sugar in the saturated sugar solution will continue to dissolve at the same rate as some is precipitated. In the unopened bottle of sparkling mineral water, the dissolved carbon dioxide will be in equilibrium with the gaseous carbon dioxide. In all of the other situations, the reverse reaction is not occurring at the same rate as the forward reaction, because the systems are open.
7.  $130 \text{ kJ mol}^{-1}$ . The activation energy is the difference between  $390 \text{ kJ mol}^{-1}$  and  $260 \text{ kJ mol}^{-1}$ , which is  $130 \text{ kJ mol}^{-1}$ .
8.
  - a Chemical equilibrium is 'dynamic' because both forward and reverse reactions occur at the same rate. An equilibrium develops between water vapour and water when wet clothes are in a sealed bag, with water evaporating as rapidly as water vapour condenses, so the clothes remain wet.
  - b When the bag is opened, water vapour escapes and the rate of evaporation exceeds the rate of condensation. The system is not in equilibrium and the clothes dry.
9.
  - a  $0.07 \text{ mol L}^{-1}$
  - b  $0 \text{ mol L}^{-1}$
  - c  $0.03 \text{ mol L}^{-1}$
  - d  $0.08 \text{ mol L}^{-1}$
  - e 0.04 moles
  - f The horizontal regions of the graph indicate that there is no change in concentration of  $\text{NO}_2$  or  $\text{N}_2\text{O}_4$ . This is when the system is at equilibrium and both the forward and reverse reactions are occurring at the same rate.
  - g 6 seconds
  - h As the reaction proceeds, the intensity of the brown colour will increase as the concentration of  $\text{NO}_2$  increases until equilibrium is reached. Once equilibrium is reached, the colour of the reaction mixture will remain constant.

# Reaction Rates An Equilibrium Answers

**CL Gary**



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1. Understanding the eBook Reaction Mrates An Equilibrium Answers
  - The Rise of Digital Reading Reaction Mrates An Equilibrium Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Mrates An Equilibrium Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reaction Mrates An Equilibrium Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Mrates An Equilibrium Answers
  - Personalized Recommendations

- Reaction Mrates An Equilibrium Answers User Reviews and Ratings
- Reaction Mrates An Equilibrium Answers and Bestseller Lists
- 5. Accessing Reaction Mrates An Equilibrium Answers Free and Paid eBooks
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  - Reaction Mrates An Equilibrium Answers Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
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  - Reaction Mrates An Equilibrium Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Reaction Mrates An Equilibrium Answers
  - Highlighting and Note-Taking Reaction Mrates An Equilibrium Answers
  - Interactive Elements Reaction Mrates An Equilibrium Answers
- 8. Staying Engaged with Reaction Mrates An Equilibrium Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Reaction Mrates An Equilibrium Answers
- 9. Balancing eBooks and Physical Books Reaction Mrates An Equilibrium Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Reaction Mrates An Equilibrium Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Mrates An Equilibrium Answers
  - Setting Reading Goals Reaction Mrates An Equilibrium Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Mrates An Equilibrium Answers
  - Fact-Checking eBook Content of Reaction Mrates An Equilibrium Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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