

EXPERIMENT 9 **CHEMICAL REACTIONS: STOICHIOMETRY** **LABORATORY REPORT**

Your Name _____ TA's Name _____ Lab Section 1411. _____ Date _____

You must read pages 69 and 70 to complete the following calculations.

	Trial #1	Trial #2	Trial #3
Vol. 0.100 M K_2CrO_4 solution	<u>12.00 mL</u>	<u>12.00 mL</u>	<u>12.00 mL</u>
mmole K_2CrO_4	<u>1.20</u>		
Vol. 0.100 M $Pb(NO_3)_2$ solution	<u>10.00 mL</u>	<u>12.00 mL</u>	<u>14.00 mL</u>
mmole $Pb(NO_3)_2$	<u>1.0</u>		
mass product + paper	<u>0.579 g</u>	<u>0.654 g</u>	<u>0.639 g</u>
mass dry filter paper	<u>0.256 g</u>	<u>0.266 g</u>	<u>0.251 g</u>
mass $PbCrO_4$ collected precip.	<u>.323</u>		
mmole $PbCrO_4$	<u>1.0 mmole</u>	<u>1.2</u>	

CALCULATIONS AND QUESTIONS

(USE THE SPACE ON THE BACK OF THIS SHEET IF NECESSARY)

1. In each trial determine which reagent was the limiting reagent? Give a reason, determine the results of your experiment, which proves your conclusion.



2. Calculate the number of millimoles of K_2CrO_4 in 10.55 mL of a 0.120 M solution. Report answer with 3 significant figures (SF).

Reaction Stoichiometry Lab Answers

AW Rasmussen



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