

CHAPTER 13

Rates of Reaction

■ SOLUTIONS TO EXERCISES

Acetaminophen is a substance that gives analgesic results in the treatment of people who will not tolerate aspirin, and the two agents together is analgesic. The high amount of the analgesic the most common people tolerate in the treatment of individuals who are very sensitive to the analgesic effect of aspirin, only the low amount of the analgesic.

- 13.1. Rate of formation of H_2O is $1/2$ of the rate. Rate of loss of O_2 is $1/2$ of the rate. Rate of loss of H_2O_2 is the same as the rate of formation of H_2O and O_2 .

$$\frac{1}{2} \frac{\Delta[\text{H}_2\text{O}]}{\Delta t} = \frac{1}{2} \frac{\Delta[\text{O}_2]}{\Delta t} = - \frac{1}{2} \frac{\Delta[\text{H}_2\text{O}_2]}{\Delta t}$$

- 13.2. Rate $= \frac{100}{1000} = 0.1$ mol/L-s. Rate of formation of H_2O is 0.1 mol/L-s. Rate of loss of O_2 is 0.1 mol/L-s.

- 13.3. The rate of loss of O_2 is 0.1 mol/L-s. The rate of loss of H_2O_2 is 0.1 mol/L-s. The rate of loss of H_2O is 0.1 mol/L-s.

- 13.4. The rate of loss of O_2 is 0.1 mol/L-s. The rate of loss of H_2O_2 is 0.1 mol/L-s. The rate of loss of H_2O is 0.1 mol/L-s.

$$\text{Rate} = \frac{100}{1000} = 0.1$$

The rate of loss of O_2 is 0.1 mol/L-s. The rate of loss of H_2O_2 is 0.1 mol/L-s. The rate of loss of H_2O is 0.1 mol/L-s.

$$\text{Rate} = \frac{100}{1000} = 0.1$$

- 13.5. The rate of loss of O_2 is 0.1 mol/L-s. The rate of loss of H_2O_2 is 0.1 mol/L-s. The rate of loss of H_2O is 0.1 mol/L-s.

$$\text{Rate} = \frac{100}{1000} = 0.1$$

The rate of loss of O_2 is 0.1 mol/L-s.

$$\text{Rate} = \frac{100}{1000} = 0.1$$

The rate of loss of O_2 is 0.1 mol/L-s.

$$\text{Rate} = \frac{100}{1000} = 0.1$$

Reaction Rates Assessment Solutions Manual

Kathleen Armour



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