#### Equilibrium Practice Problems

 Write the equilibrium expression for each of the following reactions:  $N_{2,(g)} + 3 H_{2,(g)} \leftrightarrow 2 NH_{3,(g)}$ 

$$K = \frac{[NH_3]^2}{[N_2][H_2]^3}$$

$$I_{2(s)} + CI_{2(g)} \leftrightarrow 2ICI_{(g)}$$

$$K = \underline{[ICI]^2}$$

$$NO_{2(q)} \leftrightarrow NO_{(q)} + V_2 O_{2(q)}$$

$$K = [NO][O_2]^{1/2}$$
  
 $[NO_2]$ 

The dissociation of acetic acid, CH<sub>3</sub>COOH, has an equilibrium constant at 25°C of 1.8 x 10°5. The reaction is

$$CH_3COOH_{(aq)} \leftrightarrow CH_3COO^-_{(aq)} + H^+_{(aq)}$$

If the equilibrium concentration of CH<sub>2</sub>COOH is 0.46 moles in 0.500 L of water and that of CH<sub>3</sub>COO is 8.1 x 10<sup>-3</sup> moles in the same 0.500 L, calculate [H+] for the reaction.

$$K = [CH_3COO+][H+]$$

 $K = [CH_3COO^*][H^*]$  Plug in known values and solve.

$$1.8 \times 10^{-5} = [8.1 \times 10^{-3} \text{ moles/0.500 L}] [H^+]$$
  
[0.46 moles/0.500 L]

$$[H^+] = 1.0 \times 10^{-3} M$$

Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

		catalyst	pressure	temperature	concentration
a.		presence increase	direct relationship	direct relationship	direct relationship
ь.	posn	nothing	based on moles	based on enthalpy of rxn	inverse relationship

# **Reaction Rates And Equilibrium Practice Problems**

**H Kauffman** 

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