

# Periodic and Groups: Valence Electrons: Mastery

## Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group					18A	
	1A	2A	13A	14A	15A	16A	17A	18A	
1	H 1H	He 2He							
2	Li 3Li	Be 4Be							
3	Na 11Na	Mg 12Mg	Transition metals in between these rows						
4	K 19K	Ca 20Ca							
5	Rb 37Rb	Sr 38Sr							
6	Cs 55Cs	Ba 56Ba							
7	Fr 87Fr	Ra 88Ra							

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Group 13A and Period 2?

What group and period is Helium in?

What element is in Group 2, and Period 16A? Element 2 is in Group 18A is Helium (He). The element is oxygen.

What element is in Group 2A, and Period 2?

What element Helium is?

## Valence Electrons

Valence Electrons are the most important electrons in an atom. Each period contains has the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Group 17



7 valence electrons

Group 18, including 18A and 8A, contains noble gases. How many valence electrons do they have?

How many electrons?

Hydrogen and Helium are both in Group 1A-2A, but 1 has 1 valence electron, 2 has 2 valence electrons.

**Group 18A:** - Shows you more clearly that there is 8A shell of electrons. Elements are in Group 18 and 8A are noble gases. Helium, Neon, Argon, Krypton, Xenon, and Radon are noble gases. "He 2, Ne 10, Ar 18, Kr 36, Xe 54, Rn 86" shows Group 18A has a full outer shell of valence electrons. 18A noble gases have full outer shells of 8 valence electrons.

# Periods And Groups Valence Electrons Masses Answers

**Tobias Bleicker**



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