

PERIODIC CLASSIFICATION OF ELEMENT

DOBERNIER TRIADS:-

- Arranged elements in the group of 3.
- Atomic mass of middle element is average of atomic mass of other two elements.



Drawbacks:-

- as known elements were not grouped

Lothar Meyer's Curves:-



- Alkali metals → top position
- Alkaline earth metals → descending
- Halogen → ascending position.

MODERN PERIODIC TABLE:-

- long form of periodic table
- Based on Moseley law: $\sqrt{\lambda} = a - b(Z - c)$
- Periodic properties are periodic function of their atomic number.

Periodicity - Repetition of similar properties of elements after regular interval of time.

Cause of periodicity - due to similar e^- configuration.

Transition elements; they have incompletely filled d-orbitals in their ground state.

Zn, Cd, Hg → are not transition elements because of completely filled d-orbitals. (pseudo transition elements).

Representative elements - Combined elements of s and p block are known as representative elements.

NEWLAND'S OCTAVE RULE:-

- When elements are arranged in increasing order of atomic mass, then every 8th element standing from first will show same property.



Drawbacks:-

1. Valid upto calcium only
2. No position of inert gases
3. Not worked on heavier elements.

MENDLEEV'S PERIODIC

- Periodic property of elements are periodic function of their atomic masses.

Merits:-

- Simplified study
- Some vacant space was left for the elements which were not known at that time.

Eka aluminium → Ga
Eka silicon → Ge

- Corrected atomic mass of element

Demerits:-

- (a) Position of hydrogen was not fixed
- (b) Could not explain periodicity.

MODERN PERIODIC TABLE:-

18 groups and 7 periods.

4-Blocks

s-Block = G_1 and $G_2 = ns^1-2$

p-Block = $G_{13} - G_{18} = ns^2 np^1-6$

d-Block = $G_3 - G_{10} = (n-1)d^1-10 ns^2$

f-Block = lanthanoid = $2 = 5f^1-14$

actinoids = $2 = 5f^1-14$

Diagonal Relation;

Li → Be → B → similar
Na → Mg → Al → change ratio
size

Nomenclature;

0 → nil	5 → pent	Suffix = turn
1 → un	6 → hex	100 unnilpentium
2 → bi	7 → sept	
3 → tri	8 → oct	
4 → quad	9 → enn	

EFFECTIVE NUCLEAR CHARGE

$Z_{eff} = Z - \sigma$

Trends - Period → $Z_{eff} \uparrow$

Group → Z_{eff} almost remains same. or $Z_{eff} \uparrow \downarrow$

Shielding → $s > p > d > f$

→ more +ve charge, more is Z_{eff}

→ more -ve charge, less is Z_{eff} .

IONISATION ENERGY

→ Energy required to remove an electron from an isolated gaseous atom.

Mg → $M^{+}_{(g)} + e^-$ IE₁
 $M^{+}_{(g)} \rightarrow M^{2+}_{(g)} + e^-$ IE₂

IE = IE₁ < IE₂ < IE₃ ...

Trends - $Z_{eff} \propto IE$

Period - $Z_{eff} \uparrow = IE \uparrow$

Group - $Z_{eff} \downarrow = IE \downarrow$

$G_2 = B > Al > Ga > In$

$G_{14} = C > Si > Ge > Pb > Sn$

Period 2:-

$Li < Be < B < C < N < O < F < Ne$

ATOMIC RADIUS:-

- Distance b/w nucleus and electron of outermost shell.
- But it's impossible to find exact atomic radii.

Reason-1 = atoms too small

2 = atom doesn't have any boundary

Covalent Radius - Half of the internuclear distance b/w 2 atoms bonded by covalent bond.



Metallic Radius - Half of the internuclear distance b/w 2 atoms bonded by a metallic bond.



Van der Waals Radius - Half of the internuclear distance b/w 2 nearest molecules in non-bonded state.



Trends; Atomic Radius $\propto 1/Z_{eff}$
Period → $Z_{eff} \uparrow = \text{Atomic Radius} \downarrow$
Group → $Z_{eff} \downarrow = \text{Atomic Radius} \uparrow$

Ionic Radius + Radius < Neutral Radius

ELECTRON GAIN ENTHALPY:-

→ Energy required or released when an e^- is added to outermost orbit of an isolated gaseous atom. It is known as Electron gain enthalpy.



Note - In majority cases $\Delta_{eg}H_2$ is -ve (exothermic)

Exception = B, N } $\Delta_{eg}H_2 = +ve$
Noble gas } due to stable configuration

Note - Successive electron gain enthalpy is always +ve.

Trends; $\Delta_{eg}H \propto Z_{eff}$

Period - $Z_{eff} \uparrow \Rightarrow \Delta_{eg}H \uparrow$

Group - $Z_{eff} \downarrow \Rightarrow \Delta_{eg}H \downarrow$

$G_2 = S > Se > Te > Po > O$
→ due to small size Reason!
→ e^-e^- repulsion.

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